200 Ways to Pass the Chemistry Physical Setting Regents Exam

- Protons are positively charged (*).
- Neutrons have no charge.
- Electrons are small and are negatively charged (-).
- Protons & neutrons are in an atom's nucleus (nucleons).
- Electrons are found in "clouds" (orbitals) around an atom's nucleus.
- The mass number is equal to an atom's number of protons and neutrons added together.
- The atomic number is equal to the number of protons in the nucleus of an atom.
- The number of neutrons = mass number atomic number.
- Isotopes are atoms with equal numbers of protons, but differ in their neutron numbers.
- Cations are positive (+) ions and form when a neutral atom loses electrons. They
 are smaller than their parent atom.
- Anions are negative ions and form when a neutral atom gains electrons. They are larger than their parent atom.
- Emest Rutherford's gold foil experiment showed that an atom is mostly empty space with a small, dense, positively-charged nucleus.
- J.J. Thompson discovered the electron and developed the "plum-pudding" model of the atom.

Positive & negative particles spread throughout entire atom.

- 14. Dalton's model of the atom was a solid sphere of matter that was uniform throughout.
- 15. The Bohr Model of the atom placed electrons in "planet-like" orbits around the nucleus of an atom.
- 16. The current, wave-mechanical model of the atom has electrons in "clouds" (orbitals) around the nucleus.
- 17. USE THE REFERENCE TABLESIII
- "STP" means "Standard Temperature and Pressure." (273 Kelvin & 1 atm)
- Electrons emit energy as light when they jump from higher energy levels back down to lower (ground state) energy levels. Bright line spectra are produced.
- Elements are pure substances composed of only one kind of atom.
- Binary compounds are substances made up of only two kinds of atoms. (examples: H₂O, NH₃, CO₂)
- 22. Distomic molecules are elements that form two atom molecules in their natural form at STP. Remember the phrase "BrINCIHOF" (Br₂, I₂, N₂, CL₂, H₂, O₂, F₂)
- Use this diagram to help determine the number of significant figures in a measured value...

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