

4.7 - REVIEW – Unit 4 – Chem. Compds & Bonding

1. Atoms bond to become **STABLE** which is to say it will gain a stable and full **OCTET**.
2. Ionic bonding occurs between **METALS** and **NONMETALS**. They are held together due to **OPPOSITE** charges attracting.
3. In general, **METALS** are more likely to be oxidized because they have **FEWER** valence electrons while **NONMETALS** are more likely to be reduced because they have **MORE** valence electrons.
4. When calcium reacts with bromine, the calcium becomes **OXIDIZED** (oxidized, reduced) and the bromine becomes **REDUCED** (oxidized, reduced). The oxidizing agent is **BROMINE**, and the reducing agent is **CALCIUM**.
5. A metallic bond is the attraction between the **NUCLEUS** and the **NEGATIVE** sea of **ELECTRONS**. Which occurs between two **METALS**.
6. When a transition metal forms a bond, it loses electrons from **THE S-SHELL**. For example, Zn, with electron configuration **[Ar] 4s² 3d¹⁰** will form an ion with a charge of **+2**. The ion's configuration will be **[Ar] 4s⁰ 3d¹⁰**.
7. Write the equation for the formation of the following ions (elements becoming stable).



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 - The Rise of Digital Reading Chemistry Unit 4 Concept Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 4 Concept Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Unit 4 Concept Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 4 Concept Review Answers

- Personalized Recommendations
 - Chemistry Unit 4 Concept Review Answers User Reviews and Ratings
 - Chemistry Unit 4 Concept Review Answers and Bestseller Lists
5. Accessing Chemistry Unit 4 Concept Review Answers Free and Paid eBooks
 - Chemistry Unit 4 Concept Review Answers Public Domain eBooks
 - Chemistry Unit 4 Concept Review Answers eBook Subscription Services
 - Chemistry Unit 4 Concept Review Answers Budget-Friendly Options
 6. Navigating Chemistry Unit 4 Concept Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Unit 4 Concept Review Answers Compatibility with Devices
 - Chemistry Unit 4 Concept Review Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Unit 4 Concept Review Answers
 - Highlighting and Note-Taking Chemistry Unit 4 Concept Review Answers
 - Interactive Elements Chemistry Unit 4 Concept Review Answers
 8. Staying Engaged with Chemistry Unit 4 Concept Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Unit 4 Concept Review Answers
 9. Balancing eBooks and Physical Books Chemistry Unit 4 Concept Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Unit 4 Concept Review Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Unit 4 Concept Review Answers
 - Setting Reading Goals Chemistry Unit 4 Concept Review Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Unit 4 Concept Review Answers

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 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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