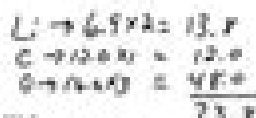


Name: _____

Date: _____



11. a. How many moles are in 1256 g of
- Li_2CO_3
- ?

$$1256 \text{ g Li}_2\text{CO}_3 \times \frac{1 \text{ mol Li}_2\text{CO}_3}{77.8 \text{ g Li}_2\text{CO}_3} = \boxed{17.0 \text{ mol}}$$

- b. If the moles from part a are added to 2.5 L of water, what is the molarity?

$$M = \frac{\text{mol}}{L} = \frac{17.0}{2.5} = \boxed{6.8 \text{ M}}$$

12. a. What is the mass of 5.9 moles of
- MgCl_2
- ?

$$\begin{aligned} \text{Mg} &\rightarrow 24.3 \times 1 = 24.3 \\ \text{Cl} &\rightarrow 35.5 \times 2 = 71.0 \\ \text{O} &\rightarrow 16.0 \times 2 = 32.0 \\ &\quad \underline{127.3} \end{aligned}$$

$$5.9 \text{ mol} \times \frac{127.3 \text{ g}}{1 \text{ mol}} = \boxed{750.97 \text{ g}}$$

- b. What would the density (in
- g/mL
-) of the substance from part a be if it has a volume of 199.2 mL?

$$D = \frac{m}{V} = \frac{750.97 \text{ g}}{199.2 \text{ mL}} = \boxed{3.77 \text{ g/mL}}$$

13. If 2.5 moles of hydrochloric acid react, how many liters of hydrogen gas at standard conditions are formed?



$$2.5 \text{ mol HCl} \times \frac{1 \text{ mol H}_2}{2 \text{ mol HCl}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = \boxed{28 \text{ L}}$$

14. The reaction in which sulfur and chlorine react to make sulfur tetrachloride is given below. How many moles of sulfur tetrachloride are produced by the reaction if 82.5 moles of chlorine are used?

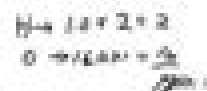
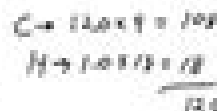
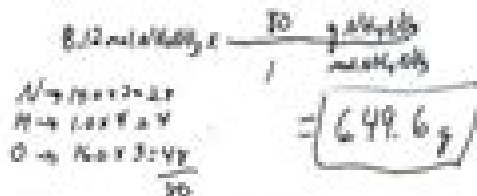
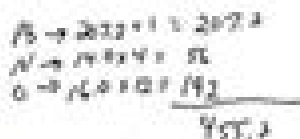


$$82.5 \text{ mol Cl}_2 \times \frac{8 \text{ mol SCl}_4}{16 \text{ mol Cl}_2} = \boxed{41.25 \text{ mol SCl}_4}$$

15. What mass of potassium nitrate will be produced if 102.2 g of lead (Pb) nitrate is reacted with excess potassium phosphate according to the reaction listed below?



$$102.2 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.2 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol Pb}_3(\text{PO}_4)_2}{3 \text{ mol Pb(NO}_3)_2} \times \frac{326 \text{ g Pb}_3(\text{PO}_4)_2}{1 \text{ mol Pb}_3(\text{PO}_4)_2} = \boxed{64.6 \text{ g}}$$



Name: _____

16. Balance these with oxygen according to the following reactions.



- a. If 301.2 g of
- C_2H_6
- was used, how many grams of
- H_2O
- could be produced?

$$301.2 \text{ g} \times \frac{1 \text{ mol C}_2\text{H}_6}{66.1 \text{ g C}_2\text{H}_6} \times \frac{6 \text{ mol H}_2\text{O}}{2 \text{ mol C}_2\text{H}_6} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{161.2 \text{ g H}_2\text{O}}$$

$$21.51 \text{ mol H}_2\text{O} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{387.2 \text{ g H}_2\text{O}}$$

- b. If 690.5 L of
- O_2
- under normal conditions was used, how many grams of
- H_2O
- could be produced?

$$690.5 \text{ L O}_2 \times \frac{1 \text{ mol O}_2}{22.4 \text{ L O}_2} \times \frac{6 \text{ mol H}_2\text{O}}{2 \text{ mol O}_2} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{369 \text{ g H}_2\text{O}}$$

$$20.5 \text{ mol H}_2\text{O} \times \frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{369 \text{ g H}_2\text{O}}$$

Also Theoretical for 20.5 (10.25 mol O_2)
It makes less

- c. Using the information from part a and b, what was the limiting reagent?



- d. Using the amount produced by your limiting reagent as your theoretical yield, what would the percentage yield be if your experimental yield was 277.4 g of
- H_2O
- ?

$$\frac{\text{Exp. yield}}{\text{Theoretical}} \times 100 = \frac{277.4}{369} \times 100 = \boxed{75.2\%}$$

- e. Using the amount produced by your limiting reagent as your theoretical yield, what would the percentage yield be if your experimental yield was 277.4 g of
- H_2O
- ?

$$\frac{\text{Exp. yield}}{\text{Theoretical}} \times 100 = \frac{277.4}{369} \times 100 = \boxed{75.2\%}$$

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L Darling-Hammond



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