

## Unit 3 Worksheet 3: Mole relationships

Four marks of 1 for questions 1-4 (total 4 marks)

a. Write the balanced chemical equation

b. Calculate volume of gas at STP when 100g of sodium reacts with sufficient water

c. Use coefficients of the balanced equation to determine mole ratio

d. Which unit are you calculating it?

11. Hydrogen sulfide gas, which smells like rotten eggs, has been used in the past to produce sodium hydride used in rocket. How many moles of hydrogen gas would be needed to completely form 1 mole of hydrosulfide as follows?



12. Ammonia,  $\text{NH}_3$ , is formed in the air by reacting hydrogen and nitrogen gases at high temperatures and pressures. How many moles of ammonia can be made from 1.00 mole of nitrogen gas?



13. Propane,  $\text{C}_3\text{H}_8$ , burns in air to form carbon dioxide and water. What moles of carbon dioxide are formed, how many moles of propane were burned?



# Chemistry Unit 8 Test B Percent Yield

**New York State College of Agriculture,  
Ithaca**



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