C2 - THE PERIODIC TABLE 2.6 The transition metals

GCSE Chemistry Aiming for Grade 4

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GCSE Chemistry A

Where are the transition metals on the periodic table? In the centre block

Sive the names of five transition metals

Bold, Silver, Aluminium, Dran, Cooper

List the physical properties of the transition maters

- 1. Seed conductors of heat
- Z. Seed conductors of electricity
- 3. High densities
- 4. High matting points
- 5. Hard old Strole

Compared to Group 1 metals how do the following properties compore?

Dengity:

Much higher than group I, none of them will fleet an mater like lithium, sedium and patezalum.

Multing beint.

Much higher melting points except for mercury

Building point.

Balling points are extremely high compared to the allkall martale.

Hardward.

Alkell metals can be cut with a knife and are sain. Trensition metals one very hard.

Street, the

Transition metals are very strong and can be used for buildings and machines due to their strength

Why is mercury not typical as a transition metal?

It is a flouid of room temperature whereas all the others are solid with high malting paints.



Sive the names of the following transition metal: compounds.

Fair Onli Lead Children

NECO4 -Nickel Corbonete Copper Civida

the state of the s

Authority. Silver Mitteste

Why is copper used for elec-It is a very good conductor other metals. It is also very drawn into wines.

Steel is made from Iron one make winders on bridges? It is very strong and hard m weight and high forces.

Titonium is used for aircraft Titlenium has a very high me: ideal for very hat engine per

State the colours of the con of transition metals.

Copper sulphore blue due to

Nickel Carbonate pole green

Chromium oxide dork press

Manganese Chloride pole pin

Componed with Group 1 elema metals have different proper Transition mateix can form a their reactivity is different. ex iron con form Fatt and Fat different reactivity.

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