

Solutions and Solubility Unit Test SCH3U

Name: _____

 K T C A Assessment of Learning (Product)
20 12 7 10

Multiple Choice. Circle the most correct answer. (20K)

1. When a liquid does not dissolve in another liquid, the two liquids may be described as being ____.

- A) saturated B) unsaturated C) miscible
D) immiscible E) None of the above

2. When 30 g of NaCl dissolves in 100 mL of water at a particular temperature it is considered to be:

- A) insoluble B) slightly or sparingly soluble
C) soluble D) None of the above

3. A solid is dissolved in water to form an aqueous solution. As dissolving proceeds, the temperature of the solution decreases. What can be concluded from this observation?

- A) More energy is required to break bonds in the solute than in the solvent.
B) More energy is released from bonds forming between solute and solvent than is needed to break bonds in the solute and solvent.
C) More energy is needed to break bonds in the solute and solvent than is released from bonds forming between the solute and solvent.
D) When the solute dissolves, a net amount of energy is released.

4. Which of the following usually makes a substance dissolve faster in a solvent?

- A) agitating the solution
B) increasing the particle size of the solute
C) lowering the temperature
D) decreasing the number of particles

5. Which of the following statements is false?

- A) Nonpolar liquids tend to be insoluble in polar liquids.
B) The weaker the attraction between the solute and solvent molecules, the greater the solubility.
C) Substances with similar intermolecular attractive forces tend to be soluble in one another.
D) The solubility of a gas increases in direct proportion to its partial pressure above the solution.
E) The solubility of gases in water decreases with increasing temperature

6. Most salts become more soluble in water as the:

- A) pressure is decreased B) pressure is increased
C) temperature is increased D) temperature is decreased

7. Which of the following is least likely to produce a solution?

Chemistry Unit Solutions Test Review Answers

Nicholas R. Romano



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