

Lewis Structure Worksheet

1. Draw the Lewis Structure for each of the following elements.

Element	Lewis Structure	Element	Lewis Structure	Element	Lewis Structure
Lithium	$\text{Li} \cdot$	Nitrogen	$\cdot \text{N} \cdot$	Silicon	$\cdot \text{Si} \cdot$
Beryllium	$\text{Be} \cdot$	Fluorine	$\cdot \text{F} \cdot$	Neon	$\cdot \text{Ne} \cdot$

2. Draw the Lewis Dot Diagram for each of the following elements. Determine how many electrons the atom has to gain or lose to become stable (fill or lose outer orbit to make 8). Determine the charge that the ion will have (recall: gain electrons = "-" charge, lose electrons = "+" charge).

Element	Lewis Structure	Needed for Stability	Charge on Ion
Sodium	$\text{Na} \cdot$	Lose 1	+1
Oxygen	$\cdot \text{O} \cdot$	Gain 2	-2
Aluminum	$\text{Al} \cdot$	Lose 3	+3
Helium	$\text{He} \cdot$	N/A	N/A
Potassium	$\text{K} \cdot$	Lose 1	+1
Carbon	$\cdot \text{C} \cdot$	4	Does not form an ion

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