

Chemistry Notes for class 12 Chapter 2 Solutions

Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution.

Solute and Solvent

In a binary solution, solvent is the component which is present in large quantity while the other component is known as solute.

Classification of Solutions

(A) Following types of solutions are seen on the basis of physical state of solute and solvent.

S.No.	Solute	Solvent	Examples
Solid solutions			
1	Solid	Solid	Alloys
2	Liquid	Solid	Hydrated salts. Amalgam of Hg with Na
3	Gas	Solid	Dissolved gases in mineral
Liquid solutions			
4	Solid	Liquid	Salt/sugar solution in water
5	Liquid	Liquid	Alcohol in water
6	Gas	Liquid	Aerated drinks, O_2 in water
Gaseous solutions			
7	Solid	Gas	Iodine vapour in air
8	Liquid	Gas	Water vapour in air
9	Gas	Gas	Air ($O_2 + N_2$)

[if water is used as a solvent, the solution is called aqueous solution and if not, the solution is called non-aqueous solution.]

(B) Depending upon the amount of solute dissolved in a solvent we have the following types of solutions:

(i) **Unsaturated solution** A solution in which more solute can be dissolved without raising temperature is called an unsaturated solution.

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Acronyms, Initialisms & Abbreviations Dictionary Gale Research Company, 1997

Books in Print Supplement, 1977 *Chemistry* Laurel Dingrando, McGraw-Hill Staff, Glencoe/McGraw-Hill, 2001-04-01

Modern Chemistry Holt, Rinehart and Winston Staff, 2006-01 **Student Study Guide/Solutions Manual to accompany General, Organic & Biological Chemistry** Erin Smith Berk, Janice

Smith, 2012-02-10 Each chapter of the Student Study Guide begins with a chapter review tied to the chapter goals in the text Next Sample problems are supplied and stepped out through the solution for each type of problem covered in the chapter A Self Test serves up fill in the blank exercises to assess learning with answers supplied at the end of the chapter Finally chapters end with the solutions for all of the in chapter problems as well as for the odd numbered end of chapter problems

Study Guide and Solutions Manual to accompany Basic Concepts of Chemistry, 9th Edition Leo J. Malone, 2011-12-07 This is the Study Guide and Solutions Manual to accompany Malone's Basic Concepts of Chemistry

Student Study Guide/Solutions Manual to accompany General, Organic & Biological Chemistry Janice Gorzynski Smith,

Dr., 2015-08-06 The Student Solutions Manual prepared by Erin R Smith and Janice Gorzynski Smith begins each chapter with a detailed chapter review that is organized around the chapter goals and key concepts The Problem Solving section provides a number of examples for solving each type of problem essential to that chapter The Self Test section of each chapter quizzes chapter highlights with answers provided Finally each chapter ends with the solutions to all in chapter problems as well as the solutions to all odd numbered end of chapter problems

Assessment of Chemistry Association for Institutional Research, Alexis Collier, 2010-11

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with hcl as 57.9 kJ mol⁻¹

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