

## Chemistry: Percent Yield

Directions: Solve each of the following problems. Show your work, including proper units, to earn full credit.

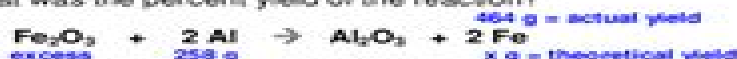
1. "Slaked lime,"  $\text{Ca(OH)}_2$ , is produced when water reacts with "quick lime,"  $\text{CaO}$ . If you start with 2 400 g of quick lime, add excess water, and produce 2 060 g of slaked lime, what is the percent yield of the reaction?



$$\text{x g Ca(OH)}_2 = 2400 \text{ g CaO} \left( \frac{1 \text{ mol CaO}}{56 \text{ g CaO}} \right) \left( \frac{1 \text{ mol Ca(OH)}_2}{1 \text{ mol CaO}} \right) \left( \frac{74 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \right) = 3171 \text{ g Ca(OH)}_2$$

$$\% \text{ Yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \Rightarrow \% \text{ Yield} = \frac{2060 \text{ g Ca(OH)}_2}{3171 \text{ g Ca(OH)}_2} \times 100 \Rightarrow 65\%$$

2. Some underwater welding is done via the thermite reaction, in which rust ( $\text{Fe}_2\text{O}_3$ ) reacts with aluminum to produce iron and aluminum oxide ( $\text{Al}_2\text{O}_3$ ). In one such reaction, 258 g of aluminum and excess rust produced 464 g of iron. What was the percent yield of the reaction?



$$\text{x g Fe} = 258 \text{ g Al} \left( \frac{1 \text{ mol Al}}{27 \text{ g Al}} \right) \left( \frac{2 \text{ mol Fe}}{2 \text{ mol Al}} \right) \left( \frac{55.8 \text{ g Fe}}{1 \text{ mol Fe}} \right) = 533 \text{ g Fe}$$

$$\% \text{ Yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \Rightarrow \% \text{ Yield} = \frac{464 \text{ g Fe}}{533 \text{ g Fe}} \times 100 \Rightarrow 87\%$$

3. Use the balanced equation to find out how many liters of sulfur dioxide are actually produced at STP if  $1.5 \times 10^{27}$  molecules of zinc sulfide are reacted with excess oxygen and the percent yield is 75%.



$$\begin{aligned} \text{x L SO}_2 &= 1.5 \times 10^{27} \text{ molecules ZnS} \left( \frac{1 \text{ mol ZnS}}{6.02 \times 10^{23} \text{ molecules ZnS}} \right) \left( \frac{2 \text{ mol SO}_2}{2 \text{ mol ZnS}} \right) \left( \frac{22.4 \text{ L SO}_2}{1 \text{ mol SO}_2} \right) \\ &= 5.58 \times 10^4 \text{ L SO}_2 \leftarrow \text{theoretical yield} \end{aligned}$$

$$\% \text{ Yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \rightarrow 0.75 = \frac{\text{X L SO}_2}{5.58 \times 10^4 \text{ L SO}_2} \rightarrow 4.19 \times 10^4 \text{ L SO}_2$$

4. The Haber process is the conversion of nitrogen and hydrogen at high pressure into ammonia, as follows:



If you must produce 700 g of ammonia, what mass of nitrogen should you use in the reaction, assuming that the percent yield of this reaction is 70%?

$$\% \text{ Yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \Rightarrow 0.70 = \frac{700 \text{ g NH}_3}{\text{X g NH}_3} \Rightarrow 1000 \text{ g NH}_3$$

$$\text{x g N}_2 = 1000 \text{ g NH}_3 \left( \frac{1 \text{ mol NH}_3}{17 \text{ g NH}_3} \right) \left( \frac{1 \text{ mol N}_2}{2 \text{ mol NH}_3} \right) \left( \frac{28 \text{ g N}_2}{1 \text{ mol N}_2} \right) = 824 \text{ g N}_2$$

Answers:

1. 65%

2. 87%

3.  $4.19 \times 10^4 \text{ L SO}_2$

4. 824 g  $\text{N}_2$

# Chemquest 34 Percent Yield Answer Key

**RD Boyd**



**Chemquest 34 Percent Yield Answer Key:**

The Top Books of the Year Chemquest 34 Percent Yield Answer Key The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous compelling novels captivating the hearts of readers worldwide. Lets delve into the realm of popular books, exploring the fascinating narratives that have captivated audiences this year. Chemquest 34 Percent Yield Answer Key : Colleen Hoover's "It Ends with Us" This poignant tale of love, loss, and resilience has gripped readers with its raw and emotional exploration of domestic abuse. Hoover expertly weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can triumph. Chemquest 34 Percent Yield Answer Key : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This intriguing historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids absorbing storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Discover the Magic : Delia Owens "Where the Crawdads Sing" This evocative coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens spins a tale of resilience, survival, and the transformative power of nature, entrancing readers with its evocative prose and mesmerizing setting. These top-selling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of compelling stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a exceptional and gripping novel that will keep you wondering until the very end. The novel is a warning tale about the dangers of obsession and the power of evil.

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