

## Chapter 3

1. An 18-g sample of element A combines completely with a 4-g sample of element B to form the compound AB. What is the mass of the compound formed?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_A + \text{Mass}_B = \text{Mass}_{AB}$$

$$\text{Mass}_{AB} = 18 \text{ g} + 4 \text{ g} = 22 \text{ g}$$

2. A substance breaks down into three component elements when it is heated. The mass of each component element is listed in the table below. What was the mass of the substance before it was heated?

Component	Mass (g)
A	39.10
B	54.94
C	64.00

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}} = 39.10 + 54.94 + 64.00 = 158.04 \text{ g}$$

3. Silver iodide powder has been used as an anti-septic and as an agent to seed clouds for rain. Silver iodide is 45.9% silver by mass. If you separate a 50-g sample of silver iodide into its elements, silver and iodine, how much silver would you have?

From the conservation of mass, the mass of silver recovered is equal to the mass of silver in the initial silver iodide sample. The amount of silver recovered would be  $50.0 \text{ g} \times 45.9\% = 50.0 \times 0.459 = 22.95 \text{ g} = 23 \text{ g}$ .

4. If 5 g of element A combines with 16 g of element B to form compound AB, how many grams of B are needed to form compound AB<sub>2</sub>? How many grams of B are needed to form AB<sub>3</sub>?

Compound AB<sub>2</sub> contains twice as much element B as does compound AB. Therefore,  $2 \times 6 \text{ g} = 12 \text{ g}$ .

There is 32 g of B in AB<sub>3</sub>. Compound AB<sub>3</sub> contains three times as much element B as does compound AB. Therefore,  $3 \times 16 \text{ g} = 48 \text{ g}$ . There is 48 g of B in AB<sub>3</sub>.

5. During a chemical reaction, 2.445 g of carbon reacts with 3.257 g of oxygen to form carbon monoxide gas. How many grams of carbon monoxide are formed in this reaction?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_{\text{carbon}} + \text{Mass}_{\text{oxygen}} = \text{Mass}_{\text{carbon monoxide}}$$

$$2.445 \text{ g} + 3.257 \text{ g} = 5.702 \text{ g}$$

6. Ibuprofen has the chemical formula C<sub>13</sub>H<sub>18</sub>O<sub>2</sub>. It is 75.69% carbon, 8.80% hydrogen, and 15.51% oxygen. How many mg of carbon does a 200-mg tablet of ibuprofen contain?

$$\text{Mass percentage of an element (\%)} =$$

$$\frac{\text{Mass of element}}{\text{Mass of compound}} \times 100\%$$

$$\text{Mass percentage}_{\text{carbon}} = \frac{\text{Mass}_{\text{carbon}}}{\text{Mass}_{\text{compound}}} \times 100\%$$

$$75.69\% \text{ carbon} = \frac{\text{Mass}_{\text{carbon}}}{200 \text{ mg}} \times 100\%$$

$$75.69\% \text{ carbon} \times \frac{200 \text{ mg}}{100\%} = \text{Mass}_{\text{carbon}}$$

$$\text{Mass}_{\text{carbon in the tablet}} = 151.38 \text{ mg}$$

7. During a chemical reaction, 4.032 g of hydrogen combined with oxygen to form 36.032 g of water. How many grams of oxygen reacted?

$$\text{Mass}_{\text{reactants}} = \text{Mass}_{\text{products}}$$

$$\text{Mass}_{\text{hydrogen}} + \text{Mass}_{\text{oxygen}} = \text{Mass}_{\text{water}}$$

$$4.032 \text{ g} + \text{Mass}_{\text{oxygen}} = 36.032 \text{ g}$$

$$\text{Mass}_{\text{oxygen}} = 36.032 \text{ g} - 4.032 \text{ g} = 32 \text{ g}$$

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**George W. Bluman**



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