

Name _____

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- The conjugate base of HSO_4^- is
 A) H_2SO_4 B) SO_4^{2-} C) H_3SO_4^+ D) HSO_4^+ E) OH^-
- The conjugate acid of HSO_4^- is
 A) H^+ B) H_2SO_4 C) HSO_4^+ D) SO_4^{2-} E) HSO_3^+
- The molar concentration of hydronium ion in pure water at 25°C is _____.
 A) 7.00 B) 1.0×10^{-7} C) 1.00 D) 1.0×10^{-14} E) 0.00
- The molar concentration of hydroxide ion in pure water at 25°C is _____.
 A) 1.00 B) 1.0×10^{-14} C) 1.0×10^{-7} D) 0.00 E) 7.00
- What is the pH of an aqueous solution at 25°C in which $[\text{OH}^-]$ is 0.0025 M?
 A) +2.60 B) -2.60 C) +11.40 D) -11.40 E) -2.25
- What is the pH of an aqueous solution at 25° that contains 3.98×10^{-9} M hydronium ion?
 A) 7.000 B) 9.000 C) 8.400 D) 5.600 E) 3.980
- What is the concentration (in M) of hydronium ions in a solution at 25°C with $\text{pH} = 4.282$?
 A) 4.28 B) 1.92×10^{-10} C) 1.66×10^4 D) 9.71 E) 5.22×10^{-5}
- Which solution below has the highest concentration of hydroxide ions?
 A) $\text{pH} = 3.21$ B) $\text{pH} = 9.82$ C) $\text{pH} = 7.93$ D) $\text{pH} = 12.59$ E) $\text{pH} = 7.00$
- What is the pH of a 0.015-M aqueous solution of barium hydroxide?
 A) 12.18 B) 1.52 C) 12.48 D) 1.82 E) 10.35
- Sodium hydroxide is a strong base. This means that _____.
 A) aqueous solutions of NaOH contain equal concentrations of H^+ (aq) and OH^- (aq)
 B) NaOH does not dissociate at all when it is dissolved in water
 C) NaOH dissociates completely to Na^+ (aq) and OH^- (aq) when it dissolves in water
 D) NaOH cannot be neutralized by a weak acid
 E) NaOH cannot be neutralized by ordinary means
- Of the following acids, _____ is not a strong acid.
 A) HNO_3 B) HCl C) HNO_2 D) HClO_4 E) H_2SO_4
- Of the following, _____ is a weak acid.
 A) HBr B) HF C) HClO_4 D) HCl E) HNO_3

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