

## Thermochemistry Review Part 1

**Definition Questions:** you should be able to use or apply the following terms. Be able to write complete definitions for,

energy	temperature	standard state
potential energy	heat	standard heat of formation
kinetic energy	specific heat capacity	exothermic
enthalpy	endothermic	Law of conservation of Energy
thermal	Hess's Law	isolated, closed, open system

1. The specific heat capacity of diamond is  $0.5050 \text{ J/g}^\circ\text{C}$ . How much energy is required to heat  $25.0 \text{ g}$  of diamond from  $10.5^\circ\text{C}$  to  $15.6^\circ\text{C}$ ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of  $50 \text{ g}$  and an initial temperature of  $90^\circ\text{C}$  is placed into  $100 \text{ mL}$  of water at a temperature of  $25^\circ\text{C}$ . The temperature of water rises to  $31.3^\circ\text{C}$ . Determine the specific heat capacity of aluminum.

Q lost (Al) = Q gained (Water)

$$-(mc\Delta t) = +mc \Delta t$$

$$-(50 \times c \times (31.3-90)) = (100 \times 4.18 \times (31.3-25))$$

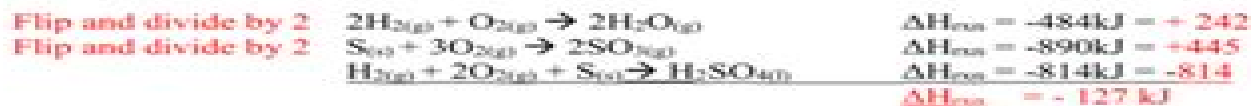
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by  $300 \text{ g}$  of methanol,  $\text{CH}_3\text{OH}$ , as it evaporates? Given: The molar heat of vaporization is  $35.3 \text{ kJ/mol}$ .

$q = ?$	$\Delta H = q/n$	$n \text{ CH}_3\text{OH} = m/\text{MM}$
$m = 300 \text{ g}$	$q = H \times n$	$n = 300 \text{ g} / 32 \text{ g/mol}$
$\Delta H = 35.3 \text{ kJ/mol}$	$q = (35.3 \text{ kJ} \times 9.375 \text{ mol})$	$n = 9.375 \text{ mol}$
	$q = 332.81 \text{ kJ}$	

4. What is the molar enthalpy of the formation of  $1 \text{ mol H}_2\text{SO}_4(\text{l})$  given the following information?



# Chemistry Thermochemistry Quiz Questions

**John Edward Feldsine**



## Chemistry Thermochemistry Quiz Questions:

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