

Thermodynamics Practice Test

Name: _____

Note - this practice test does NOT contain questions from all sections of the chapters
NChO 1999

____ 27. According to this information, and given the fact that for water, $\Delta H_f = -286 \text{ kJ mol}^{-1}$, what is ΔH_f for MgO(s) ?

| Reaction | ΔH |
|---|----------------------------|
| $\text{Mg(s)} + 2 \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$ | -467 kJ mol^{-1} |
| $\text{MgO(s)} + 2 \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{O(l)}$ | -151 kJ mol^{-1} |

- (A) -904 kJ mol^{-1}
- (B) -602 kJ mol^{-1}
- (C) -334 kJ mol^{-1}
- (D) -30 kJ mol^{-1}

NChO 1998

____ 22. Carbon reacts with oxygen according to this equation. $2\text{C(s)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO(g)}$
 $\Delta H = -220 \text{ kJ}$

Which statements are true?

1. The reaction is exothermic.
 2. The combustion of 0.50 mol of carbon produces 55 kJ of heat energy
- (A) 1 only
 - (B) 2 only
 - (C) both 1 and 2
 - (D) neither 1 nor 2

____ 25. A 1.0 g sample of substance A at 100°C is added to 100 ml of H_2O at 25°C . Using separate 100 mL portions of H_2O , the procedure is repeated with substance B and then with substance C. How will the final temperatures of the water compare?

| Substance | Specific Heat |
|-----------|---|
| A | $0.60 \text{ J g}^{-1} ^\circ\text{C}^{-1}$ |
| B | $0.40 \text{ J g}^{-1} ^\circ\text{C}^{-1}$ |
| C | $0.20 \text{ J g}^{-1} ^\circ\text{C}^{-1}$ |

- (A) $T_c > T_b > T_a$
- (B) $T_b > T_a > T_c$
- (C) $T_a > T_b > T_c$
- (D) $T_a = T_b = T_c$

____ 26. How many grams of benzene, $\text{C}_6\text{H}_6\text{(l)}$, must be burned in a bomb calorimeter to raise its temperature by 1.5°C ?
Given: The calorimeter constant is $12.59 \text{ kJ } ^\circ\text{C}^{-1}$ and the $\Delta H^\circ_{\text{combustion}}$ for $\text{C}_6\text{H}_6 = -41.9 \text{ kJ g}^{-1}$

- (A) 0.45 g
- (B) 2.8 g
- (C) 3.3 g
- (D) 8.4 g

Chemistry Thermodynamics Practice Questions

Ralph H. Petrucci, Robert K. Wismer



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1. Understanding the eBook Chemistry Thermodynamics Practice Questions
 - The Rise of Digital Reading Chemistry Thermodynamics Practice Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Thermodynamics Practice Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Thermodynamics Practice Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Thermodynamics Practice Questions
 - Personalized Recommendations
 - Chemistry Thermodynamics Practice Questions User Reviews and Ratings

- Chemistry Thermodynamics Practice Questions and Bestseller Lists
- 5. Accessing Chemistry Thermodynamics Practice Questions Free and Paid eBooks
 - Chemistry Thermodynamics Practice Questions Public Domain eBooks
 - Chemistry Thermodynamics Practice Questions eBook Subscription Services
 - Chemistry Thermodynamics Practice Questions Budget-Friendly Options
- 6. Navigating Chemistry Thermodynamics Practice Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Thermodynamics Practice Questions Compatibility with Devices
 - Chemistry Thermodynamics Practice Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Thermodynamics Practice Questions
 - Highlighting and Note-Taking Chemistry Thermodynamics Practice Questions
 - Interactive Elements Chemistry Thermodynamics Practice Questions
- 8. Staying Engaged with Chemistry Thermodynamics Practice Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Thermodynamics Practice Questions
- 9. Balancing eBooks and Physical Books Chemistry Thermodynamics Practice Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Thermodynamics Practice Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Thermodynamics Practice Questions
 - Setting Reading Goals Chemistry Thermodynamics Practice Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Thermodynamics Practice Questions
 - Fact-Checking eBook Content of Chemistry Thermodynamics Practice Questions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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