

## Chemistry Unit 7

### Worksheet 3: Adjusting to Reality - Limiting Reactant

1. Given 4 molecules of hydrogen gas and 4 molecules of oxygen gas to form water. Write the balanced equation for the reaction between hydrogen and oxygen.

Balanced Equation: \_\_\_\_\_

Make a drawing that represents the reaction container before and after the reaction.



*Before*



*After*

- \_\_\_\_\_ How many molecules of water can be produced?  
 \_\_\_\_\_ Which reactant is in excess? Why?  
 \_\_\_\_\_ How many molecules of excess reactant are there?

Construct a Before-Change-After Table for this reactant mixture:

Bal. Equation: \_\_\_\_\_  
 Before: \_\_\_\_\_  
 Change: \_\_\_\_\_  
 After: \_\_\_\_\_

According to the table you just made,

- \_\_\_\_\_ How many molecules of water can be produced?  
 \_\_\_\_\_ Which reactant is in excess? Why?  
 \_\_\_\_\_ How many molecules of excess reactant are there?

Based on your two methods of analysis above, what determines how much product can be made from a particular reactant mix?

# Chemistry Unit 7 Ws 4 Answers

**PT Brinkman**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the PT Brinkman text.

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