

Unit 2 Worksheet 3 – PVTn Problems

On each of the problems below, start with the given P, V, T, or n; then make a decision as to how a change in P, V, T, or n will affect the starting quantity, and then multiply by the appropriate factor. Draw particle diagrams of the initial and final conditions.

1. What would be the new pressure if 250 cm³ of gas at standard pressure is compressed to a volume of 150 cm³? (n and T = constant)

	P	T	V	n
Initial	1 atm	—	250 cm ³	—
Final	?	—	150 cm ³	—
Effect	↑	—	↓	—

$$P_2 = 1 \text{ atm} \times \frac{250 \text{ cm}^3}{150 \text{ cm}^3} = 1.7 \text{ atm}$$

or 1.340 atm / kg
or 1.704 Pa
or 25 psi

2. The pressure in a bicycle tire is 105 psi at 25°C in Fresno. You take the bicycle up to Huntington, where the temperature is – 5°C. What is the pressure in the tire? (V and n = constant)

	P	T	V	n
Initial	105 psi	298 K	—	—
Final	?	268 K	—	—
Effect	↓	↓	—	—

$$P_2 = 105 \text{ psi} \times \frac{268 \text{ K}}{298 \text{ K}} = 94.4 \text{ psi}$$

3. A sample of gas occupies 150 mL at 25 °C. What is its volume when the temperature is increased to 50°C? (P and n = constant)

	P	T	V	n
Initial	—	298 K	150 mL	—
Final	—	323 K	?	—
Effect	—	↑	↑	—

$$V_2 = 150 \text{ mL} \times \frac{323 \text{ K}}{298 \text{ K}} = 163 \text{ mL} \rightarrow 160 \text{ mL}$$

Chemistry Unit Pvtn Problems Answers

Frederick K. Ault



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Problems and Solutions for General Chemistry and College Chemistry, Sixth Editions by Nebergall, Holtzclaw, and Robinson Frederick K. Ault, 1980

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