

7.4NH Percent Yield (For additional information see page 317 in the book)

Practice Problems –**Work these on a Separate sheet, You will need the room.**

The theoretical yield of beryllium chloride was 10.7 grams. If the reaction actually yields 4.5 grams, what was the percent yield?

$$\% \text{ Yield} = \frac{\text{Actual yield}}{\text{Theoretical Yield}} \times 100\% = \frac{4.5 \text{ g}}{10.7 \text{ g}} \times 100\% = 42\%$$



a) A scientist began this reaction with 20 grams of lithium hydroxide and an unlimited amount of KCl. What is the theoretical yield of lithium chloride (grams)?

$$\frac{20 \text{ g LiOH}}{1} \times \frac{1 \text{ mole LiOH}}{24 \text{ g LiOH}} \times \frac{1 \text{ mole LiCl}}{1 \text{ mole LiOH}} \times \frac{42 \text{ g LiCl}}{1 \text{ mole LiCl}} = 35 \text{ g of LiCl}$$

b) The reaction actually produced 6 grams of lithium chloride. What is the percent yield?

$$\% \text{ Yield} = \frac{\text{Actual yield}}{\text{Theoretical Yield}} \times 100\% = \frac{6 \text{ g}}{35 \text{ g}} \times 100\% = 17\%$$

3. Nitrogen and hydrogen at high temperatures are converted to ammonia using the following reaction known as the Haber Process.



When 400 grams of H_2 are added to an excess amount of N_2 , 104 grams of NH_3 are formed. Calculate the percent yield? (Answer: 4.5%)

Actual Yield = 104 g of NH_3

Use 400 g of H_2 in Stoichiometry Equation to calculate the Theoretical Yield.

$$\frac{400 \text{ g H}_2}{1} \times \frac{1 \text{ mole H}_2}{2 \text{ g H}_2} \times \frac{2 \text{ moles NH}_3}{3 \text{ moles H}_2} \times \frac{17 \text{ g NH}_3}{1 \text{ mole NH}_3} = 2,267 \text{ g of NH}_3$$

$$\% \text{ Yield} = \frac{\text{Actual yield}}{\text{Theoretical Yield}} \times 100\% = \frac{104 \text{ g}}{2,266 \text{ g}} \times 100\% = 4.5\%$$

Chemistry Theoretical And Percent Yield Answers

Martha Mackin



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