

## Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H<sub>2</sub>O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<hr/> 58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<hr/> 84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<hr/> 159.6

# Chemistry Unit 5 Counting Particles Objectives Answers

## Worksheet Two

**Carolyn Chapman**



## **Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two:**

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Verma, Dr. Kulbhushan Prakash, 2006

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## **Table of Contents Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two**

1. Understanding the eBook Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
  - The Rise of Digital Reading Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
  - Personalized Recommendations
  - Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two User Reviews and Ratings
  - Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two and Bestseller Lists
5. Accessing Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two Free and Paid eBooks
  - Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two Public Domain eBooks
  - Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two eBook Subscription Services
  - Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two Budget-Friendly Options
6. Navigating Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two eBook Formats

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- ePub, PDF, MOBI, and More
  - Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two Compatibility with Devices
  - Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
    - Highlighting and Note-Taking Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
    - Interactive Elements Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
  8. Staying Engaged with Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
  9. Balancing eBooks and Physical Books Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
    - Setting Reading Goals Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
    - Fact-Checking eBook Content of Chemistry Unit 5 Counting Particles Objectives Answers Worksheet Two
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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