

11. A 70.0 g sample of limestone consists of a large percentage of calcium carbonate. The sample reacts with an excess of hydrochloric acid and 14.0 L of carbon dioxide is generated at STP. What is the percentage of calcium carbonate in the limestone? Write the balanced chemical equation for this reaction.



$$14.0\text{ L CO}_2 \left(\frac{1\text{ mol CO}_2}{22.4\text{ L CO}_2} \right) \left(\frac{1\text{ mol CaCO}_3}{1\text{ mol CO}_2} \right) \left(\frac{100.09\text{ g CaCO}_3}{1\text{ mol CaCO}_3} \right) = 62.55\text{ g CaCO}_3$$

$$\% \text{ CaCO}_3 = \frac{62.55\text{ g CaCO}_3}{70.0\text{ g limestone}} \times 100 = \boxed{89.4\% \text{ CaCO}_3}$$

2.2.5 Analyze quantitatively the composition of a substance (empirical formula, molecular formula, percent composition, anhydrides).

1. Analysis shows a compound to be, by mass, 43.8% N, 6.2% H, and 50.0% O. Which is the possible molecular formula for the substance? Assume 100 g sample

a. NH_4NO_3 $\text{N } 43.8\text{ g} \left(\frac{1\text{ mol N}}{14.01\text{ g N}} \right) = 3.13\text{ mol N}$ $\frac{50.0\text{ g O} \left(\frac{1\text{ mol O}}{16.00\text{ g O}} \right) = 3.125\text{ mol O}}{3.13\text{ mol N} \approx 1}$

b. NH_4NO_2

c. NH_4OH

d. $\text{N}_2\text{O}_4\text{H}$ $\text{H } 6.2\text{ g} \left(\frac{1\text{ mol H}}{1.008\text{ g H}} \right) = 6.15\text{ mol H}$ $\frac{6.15\text{ mol H}}{3.125\text{ mol O}} = 1.97 \approx 2$

N : H : O
1 : 2 : 1
 NH_2O empirical formula

2. A compound has an empirical formula of CH_2O and a molecular mass of 180g. What is the compound's molecular formula?

a. $\text{C}_3\text{H}_6\text{O}_3$ $\text{C } 1 \times 12.01 = 12.01$

b. $\text{C}_6\text{H}_{12}\text{O}_6$ $\text{H } 2 \times 1.008 = 2.016$ $\frac{180\text{ g}}{30.026\text{ g}} = 5.99 \approx 6$ $6(\text{CH}_2\text{O})$

c. $\text{C}_3\text{H}_4\text{O}_3$ $\text{O } 1 \times 16.00 = 16.00$

d. $\text{C}_{12}\text{H}_{20}\text{O}_{11}$ $\frac{16.00}{30.026}$ $\text{C}_6\text{H}_{12}\text{O}_6$

3. What is the percent by mass of iron in the compound Fe_2O_3 ?

a. 70%

b. 56%

c. 48%

d. 30%

Fe $2 \times 55.85 = 111.7$

O $3 \times 16.00 = 48.00$

159.70

$\frac{111.7\% \text{ mol Fe}}{159.7\% \text{ mol Fe}_2\text{O}_3} \times 100 = 69.9\%$

4. A compound consisting of 56.38% phosphorus and 43.62% oxygen has a molecular mass of 220 g/mole. What is the molecular formula of this compound? Assume 100g sample

a. PO

b. PO_2

c. P_2O_5

d. P_2O_6

$56.38\text{ g P} \left(\frac{1\text{ mol P}}{30.97\text{ g P}} \right) = 1.82\text{ mol P}$

$43.62\text{ g O} \left(\frac{1\text{ mol O}}{16.00\text{ g O}} \right) = 2.73\text{ mol O}$

$\frac{2.73\text{ mol O}}{1.82\text{ mol P}} = 1.498 \approx 1.5$

P : O
(1 : 1.5) $\times 2$ P_2O_3 empirical formula

$2(30.97) + 3(16.00) = 109.94$

$\frac{220\text{ g/mol}}{109.94} \approx 2$

5. A 10.10 g sample of barium chloride hydrate is heated in a crucible. After all of the water is driven off, 8.50 g of the anhydrous barium chloride remains in the crucible. What is the formula of the hydrate?

① 10.10g hydrate - 8.50g BaCl_2 = 1.60g H_2O

② $8.50\text{ g BaCl}_2 \left(\frac{1\text{ mol BaCl}_2}{208.23\text{ g BaCl}_2} \right) = 0.0408\text{ g BaCl}_2$

$1.60\text{ g H}_2\text{O} \left(\frac{1\text{ mol H}_2\text{O}}{18.01\text{ g H}_2\text{O}} \right) = 0.0888\text{ mol H}_2\text{O}$

③ $\frac{\text{mol H}_2\text{O}}{\text{mol BaCl}_2} = 2.04 \approx 2$

④ Answer
 $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$

Chemistry Unit 7 2 Answers

Roger Norris

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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