

Unit 7 Study Guide:

Atoms, Elements/Compounds/Mixtures, States of Matter, Physical and Chemical Properties and Changes

1. Famous scientists in the study of the atom: Name, dates, and contribution. See summary on page 203.

	Date	Contribution
Aristotle	425BC	Earth, Air, Fire, Water, Aether
Democritus	425BC	Matter is made of indivisible atoms
John Dalton	1802	Atomic Theory of Matter
Joseph Priestly	1772	Discovered Oxygen
Antoine Lavoisier	1783	1 st Chemistry Textbook, L.O.C.M.
J.J. Thomson	1897	Discovered the electron
Ernest Rutherford	1909	Discovered the proton
Niels Bohr	1913	Proposed electron energy levels
Gilbert Lewis	1916	Proposed the term 'valence'
James Chadwick	1932	Discovered the neutron

2. Using the Periodic Table to find the characteristics of an atom:

Consider Phosphorous P-15. See practice pages 211, 213, 218.

	P	Bohr Model	Lewis Model
Atomic Symbol:	P		
Atomic Number:	15		
Atomic Mass:	30.97		
Protons:	15		
Electrons:	15		
Neutrons:	16		
Group:		5	
Period:	2		
Electron Shells:	3		
Valence Electrons:	5		
Ionic State:	-3		
Type:	non metal		
Most common Isotope:	P-31		



3. Classifying matter; mixture or pure substance, homogeneous or heterogeneous, and compound or element.

Chemistry Unit 7 Study Guide

K Morrison



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Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

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