

ChemQuest 31

Using Moles with Formulas

Name: _____

Date: _____

Hour: _____

Information: Percent Composition

Sometimes it is needful to know the composition of a compound. For example, 39.3% of the mass of sodium chloride is due to sodium. The other 60.7% of the mass is from chlorine. So, in a 100 g sample of sodium chloride, there are 39.3 g of sodium and 60.7 g of chlorine. This type of data is known as **percent composition**. The percent composition tells you the percentage by mass of an element in a compound. There is a convenient formula for finding the percent composition of an element in a compound:

$$\text{percent composition of element "X"} = \frac{\text{mass of x in one mole of the compound}}{\text{mass of one mole of the compound}} \cdot 100$$

(obtained from periodic table) → (points to numerator)
 (obtained from periodic table) → (points to denominator)

Let us look at how the percent composition of calcium (Ca) in calcium chloride (CaCl_2) was determined.

$$\text{percent composition of Ca} = \frac{\text{mass of Ca in one mole of CaCl}_2}{\text{mass of one mole of CaCl}_2} \cdot 100$$

$$\text{percent composition of Ca} = \frac{40.1 \text{ g}}{111.1 \text{ g}} \cdot 100 = 36.1\%$$

from periodic table for calcium (points to 40.1 g)
 from periodic table for calcium + 2 chlorines:
 $40.1 + 2(35.5) = 111.1$ (points to 111.1 g)

As another example, consider calculating the percent composition of nitrogen in Ca_3N_2 :

$$\text{percent composition of N} = \frac{28.0 \text{ g}}{148.3 \text{ g}} \cdot 100 = 18.9\%$$

From periodic table for 2 nitrogen atoms: $2(14.0) = 28.0$ (points to 28.0 g)
 from periodic table for 3 calcium + 2 nitrogen atoms:
 $3(40.1) + 2(14.0) = 148.3$ (points to 148.3 g)

Critical Thinking Questions

1. Verify that in C_4H_{10} the percent composition of carbon is approximately 82.6%.

$$\frac{4(12.0)}{[4(12.0) + 10(1.01)]} \cdot 100 = 82.6\%$$

Note: 12.0 and 1.01 are the molar masses for carbon and hydrogen on the periodic table.

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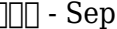
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