

## ChemQuest 48

# pH and pOH

$$\text{pH} = \underline{\hspace{2cm}}$$

$$\text{pOH} = \underline{\hspace{2cm}}$$

$$\text{pH} + \text{pOH} = \underline{\hspace{2cm}}$$

## Information: Molecular Models

Water and the hydronium ion ( $\text{H}_3\text{O}^+$ ) are shown below. The 1 positive charge which forms water is indicated above the hydrogen. The dots represent the combination of hydrogen and oxygen's valence electrons. The dots in water represent pairs of electrons. The lone pair of electrons on the oxygen atom represents the negative charge. The lone pair of electrons on the hydrogen atom represents the positive charge.

For pH calculations use 4.0. For pOH calculations use 14.0.

## Calculate: Problems/Questions

1. Calculate the pH of a solution of hydrochloric acid of  $2.5 \times 10^{-3} \text{ M}$ . What is  $\text{pH}$ ?  
 $\text{pH} = \log(2.5 \times 10^{-3}) = \underline{\hspace{2cm}}$
2. Calculate the pOH of a 0.001 M solution of hydrochloric acid of  $1.0 \times 10^{-3} \text{ M}$ . What is the  $\text{pOH}$ ?  
 $\text{pOH} = \log(1.0 \times 10^{-3}) = \underline{\hspace{2cm}}$
3. Calculate the pH of 0.01 M. What is the equivalent of hydrochloric acid?  
 $10^{-2} \times 10^{-2} = 10^{-4} = 0.0001 \text{ M}$

# Chemquest Ph And Poh Answer Key

**Didier Musso**



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