

Name: _____ Date: _____ Period: _____

Naming WS #3: Ionic Compounds with Transition Metals

Part A Write the formula for each ionic compound.

	Name	Formula
1.	Tin (II) Hydroxide	
2.	Iron (III) oxalate	
3.	Cobalt (II) bromide	
4.	Chromium (III) Chloride	
5.	Iron (II) Oxide	
6.	Mercury (II) Chloride	
7.	Tin (IV) Sulfide	
8.	Silver (I) Phosphide	
9.	Lead (IV) Iodide	

Part B Write the Name for each formula
Be sure to include the roman numeral.

	Formula	Name
1.	$\text{Fe}(\text{NO}_3)_3$	
2.	$\text{Mn}(\text{OH})_2$	
3.	$\text{Ti}(\text{NO}_3)_4$	
4.	PbS_2	
5.	CuCl	
6.	PbSO_4	
7.	ZnCl_2	
8.	Hg_2O	
9.	Cr_2O_3	

Part C: Fill in the following table. Polyatomic ions and transition metals are mixed up.

	Ionic Formula	Ionic Compound Name	Balanced Ion Pairs
1.		Sodium Chromate	
2.	CuCl		
3.		Sodium Nitrate	
4.	ZnSO_4		
5.		Potassium cyanide	
6.	Al_2O_3		
7.		Copper(II) nitrate	
8.	FePO_4		
9.		Lead(II) sulfide	
10.	NaClO		
11.		Sodium Carbonate	
12.	Na_2SeO_4		
13.		Ammonium Bromate	
14.	$\text{Au}(\text{ClO}_4)_3$		
15.		Iron(III) Sulfate	

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