

11

Electrochemistry

- When 0.1 mole of MnO_4^{2-} is oxidised, the quantity of electricity required to completely oxidise MnO_4^{2-} to MnO_4^- is
(a) 96500 C (b) 2×96500 C [2014]
(c) 9650 C (d) 96.50 C
- The weight of silver (at. wt. = 108) displaced by a quantity of electricity which displaces 5600 mL of O_2 at STP will be [2014]
(a) 5.4 g (b) 10.8 g (c) 54.0 g (d) 108.0 g
- A hydrogen gas electrode is made by dipping platinum wire in a solution of HCl of pH = 10 and by passing hydrogen gas around the platinum wire at 1 atm pressure. The oxidation potential of electrode would be [2013]
(a) 0.059 V (b) 0.59 V (c) 0.118 V (d) 1.18 V
- At 25°C molar conductance of 0.1 molar aqueous solution of ammonium hydroxide is $9.54 \Omega^{-1} \text{cm}^2 \text{mol}^{-1}$ and at infinite dilution its molar conductance is $238 \Omega^{-1} \text{cm}^2 \text{mol}^{-1}$. The degree of ionisation of ammonium hydroxide at the same concentration and temperature is [2013]
(a) 2.080 % (b) 20.800 %
(c) 4.008 % (d) 40.800 %
- A button cell used in watches functions as following [2013]
 $\text{Zn}(s) + \text{Ag}_2\text{O}(s) + \text{H}_2\text{O}(l) \rightleftharpoons 2 \text{Ag}(s) + \text{Zn}^{2+}(aq) + 2\text{OH}^-(aq)$
If half-cell potentials are
 $\text{Zn}^{2+}(aq) + 2e^- \longrightarrow \text{Zn}(s), E^\circ = -0.76 \text{ V}$
 $\text{Ag}_2\text{O}(s) + \text{H}_2\text{O}(l) + 2e^- \longrightarrow 2\text{Ag}(s) + 2\text{OH}^-(aq), E^\circ = 0.34 \text{ V}$
The cell potential will be
(a) 1.10 V (b) 0.42 V (c) 0.84 V (d) 1.34 V
- Limiting molar conductivity of NH_4OH (i.e. $\Lambda_m(\text{NH}_4\text{OH})$) is equal to [2012]
(a) $\Lambda_m(\text{NH}_4\text{Cl}) + \Lambda_m(\text{NaOH}) - \Lambda_m(\text{NaCl})$
(b) $\Lambda_m(\text{NaOH}) + \Lambda_m(\text{NaCl}) - \Lambda_m(\text{NH}_4\text{Cl})$
(c) $\Lambda_m(\text{NH}_4\text{OH}) + \Lambda_m(\text{NH}_4\text{Cl}) - \Lambda_m(\text{NaCl})$
(d) $\Lambda_m(\text{NH}_4\text{Cl}) + \Lambda_m(\text{NaOH}) - \Lambda_m(\text{NaCl})$
- Standard electrode potential of three metals X, Y and Z are -1.2V, +0.5V and -3.0V respectively. The reducing power of these metals will be [2011]
(a) $Y > X > Z$ (b) $Z > X > Y$
(c) $X > Y > Z$ (d) $Y > Z > X$
- If the E_{cell}° for a given reaction has a negative value then which of the following gives the correct relationships for the values of ΔG° and K_{eq} ? [2011]
(a) $\Delta G^\circ < 0, K_{\text{eq}} > 1$ (b) $\Delta G^\circ < 0, K_{\text{eq}} < 1$
(c) $\Delta G^\circ > 0, K_{\text{eq}} < 1$ (d) $\Delta G^\circ > 0, K_{\text{eq}} > 1$
- The electrode potentials for
 $\text{Cu}^{2+}(aq) + e^- \longrightarrow \text{Cu}^+(aq)$
and $\text{Cu}^+(aq) + e^- \longrightarrow \text{Cu}(s)$
are +0.15 V and +0.50 V respectively. The value of $E_{\text{Cu}^{2+}/\text{Cu}}^\circ$ will be [2011]
(a) 0.325 V (b) 0.650 V (c) 0.150 V (d) 0.500 V

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Table of Contents Chemsheets Electrochemistry Answers A046

1. Understanding the eBook Chemsheets Electrochemistry Answers A046
 - The Rise of Digital Reading Chemsheets Electrochemistry Answers A046
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemsheets Electrochemistry Answers A046
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemsheets Electrochemistry Answers A046
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemsheets Electrochemistry Answers A046

- Personalized Recommendations
 - Chemsheets Electrochemistry Answers A046 User Reviews and Ratings
 - Chemsheets Electrochemistry Answers A046 and Bestseller Lists
5. Accessing Chemsheets Electrochemistry Answers A046 Free and Paid eBooks
 - Chemsheets Electrochemistry Answers A046 Public Domain eBooks
 - Chemsheets Electrochemistry Answers A046 eBook Subscription Services
 - Chemsheets Electrochemistry Answers A046 Budget-Friendly Options
 6. Navigating Chemsheets Electrochemistry Answers A046 eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemsheets Electrochemistry Answers A046 Compatibility with Devices
 - Chemsheets Electrochemistry Answers A046 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemsheets Electrochemistry Answers A046
 - Highlighting and Note-Taking Chemsheets Electrochemistry Answers A046
 - Interactive Elements Chemsheets Electrochemistry Answers A046
 8. Staying Engaged with Chemsheets Electrochemistry Answers A046
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemsheets Electrochemistry Answers A046
 9. Balancing eBooks and Physical Books Chemsheets Electrochemistry Answers A046
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemsheets Electrochemistry Answers A046
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemsheets Electrochemistry Answers A046
 - Setting Reading Goals Chemsheets Electrochemistry Answers A046
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemsheets Electrochemistry Answers A046

- Fact-Checking eBook Content of Chemsheets Electrochemistry Answers A046
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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