

ChemThink: Behavior of Gases

Name _____ Per _____

1. What does the speed of atoms depend on? Describe the relationship between speed and each of these factors.

temperature: \uparrow temp = \uparrow speed

type of atom: smaller/lighter = faster

2. In a sample of one type of atom, do all the atoms move at the same speed? Why?

No, speed is related to each individual atoms energy

3. What is the equation for Pressure and what are the units for pressure?

$$\text{Pressure} = \frac{\text{Force}}{\text{Area}} \quad \text{psi}$$

4. When atoms are moving in a closed container what is creating the force on the sides of the container?

impacts (collisions) of atoms w/ sides of container

5. Explain how a change in temperature causes pressure to change and include an equation.

more collisions results in higher pressure

\downarrow
b/c molecules moving faster

$$\frac{P}{T} = \text{constant}$$

6. Draw the graph you observed of pressure vs. temperature. Is this a direct relationship or an inverse relationship?



direct

7. Fill the cause/effect relationship in the diagram of boxes you observed.



- Explain how the number of atoms present causes pressure to change and include an equation?

more atoms = more collisions, which increases pressure

$$\frac{P}{n} = \text{constant}$$

Chemthink Gases Question Answers

VM Jensen



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Gases and Gas Laws Hobsons Publishing, PLC, 1983-01-01 **Gases** , 1984

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