

# The Mole Study Guide

**Mass Composition:** What's your element's %? given after the decimal? Molar mass to obtain your answer!

10. There are 3.75 moles of  $\text{CO}_2$ .

11. There are \_\_\_\_\_ moles of \_\_\_\_\_.

12. There are \_\_\_\_\_ moles of \_\_\_\_\_.

13. There are \_\_\_\_\_ moles of \_\_\_\_\_.

14. There are 2.5 moles of \_\_\_\_\_.

15. There are \_\_\_\_\_ moles of \_\_\_\_\_.

16. There are \_\_\_\_\_ moles of \_\_\_\_\_.

17. There are \_\_\_\_\_ moles of \_\_\_\_\_.

18. There are 4.5 moles of \_\_\_\_\_.

19. There are \_\_\_\_\_ moles of \_\_\_\_\_.

20. There are \_\_\_\_\_ moles of \_\_\_\_\_.

21. There are 3.5 moles of \_\_\_\_\_.

22. There are \_\_\_\_\_ moles of \_\_\_\_\_.

23. There are \_\_\_\_\_ moles of \_\_\_\_\_.

24. There are 2.5 moles of \_\_\_\_\_.

25. There are \_\_\_\_\_ moles of \_\_\_\_\_.

26. There are \_\_\_\_\_ moles of \_\_\_\_\_.

Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

## The Mole Study Guide

Fill in the Blank boxes for the image below.



Calculate the Molar Mass for the following Compounds.

10.  $\text{H}_2\text{O}$

21.  $\text{H}_2\text{SO}_4$

Calculate the Percent Composition for each element in the following Compounds.

10.  $\text{H}_2\text{O}$

11.  $\text{H}_2\text{SO}_4$

21.  $\text{H}_2\text{SO}_4$

12.  $\text{H}_2\text{SO}_4$

13.  $\text{H}_2\text{SO}_4$

14.  $\text{H}_2\text{SO}_4$

15.  $\text{H}_2\text{SO}_4$

16.  $\text{H}_2\text{SO}_4$

17.  $\text{H}_2\text{SO}_4$

18.  $\text{H}_2\text{SO}_4$

19.  $\text{H}_2\text{SO}_4$

Write the Empirical Formula for each compound.  
(Use subscripts to show the relative ratio of atoms.)

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_2\text{H}_4\text{O}_2$	
$\text{C}_3\text{H}_6\text{O}_3$	
$\text{C}_5\text{H}_{10}\text{O}_5$	

Determine the Empirical and Molecular Formula for the compound given the information below.

10. A compound has the following percent composition: 40.00% Hg, 21.87% Cl, 3.12% N, and 34.91% O. The molar mass for Astatine is 210.0 g/mol.

a. If you have 100 g of Astatine, how many grams of Hg, Cl, and O do you have?

b. Determine the Empirical Formula for Astatine.

c. Determine the Molecular Formula for Astatine.

# Chemistry Study Guide The Mole Answers

**Alan S. Wingrove, Robert L. Caret**



## **Chemistry Study Guide The Mole Answers:**

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