

CONJUGATE ACID-BASE PAIRS

Name _____

In the exercise, Bronsted-Lowry Acids and Bases, it was shown that after an acid has given up its proton, it is capable of getting back that proton and acting as a base. Conjugate base is what is left after an acid gives up a proton. The stronger the acid, the weaker the conjugate base. The weaker the acid, the stronger the conjugate base.

Fill in the blanks in the table below.

Conjugate Pairs

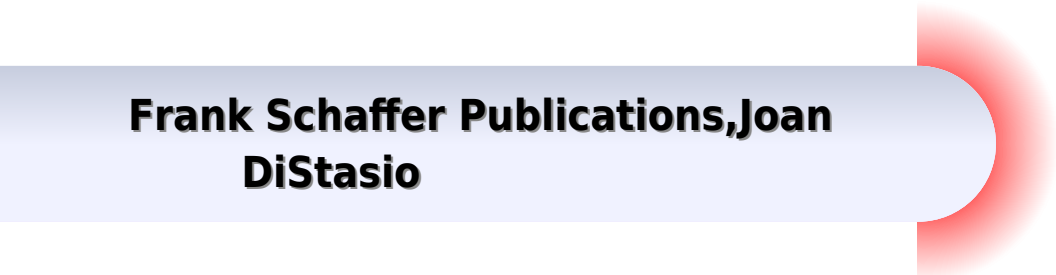
	ACID	BASE	EQUATION
1.	H_2SO_4	HSO_4^-	$\text{H}_2\text{SO}_4 \leftrightarrow \text{H}^+ + \text{HSO}_4^-$
2.	H_3PO_4		
3.		F^-	
4.		NO_3^-	
5.	H_2PO_4^-		
6.	H_2O		
7.		SO_4^{2-}	
8.	HPO_4^{2-}		
9.	NH_4^+		
10.		H_2O	

Which is a stronger base, HSO_4^- or H_2PO_4^- ? _____

Which is a weaker base, Cl^- or NO_3^- ? _____

Conjugate Acid Base Pairs Answer Key Chemistry 1 7

**Frank Schaffer Publications, Joan
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