

CHAPTER 16, Solutions *(continued)*

7. Look at Figure 16.2 on page 472. Circle the letter of each sentence that is true about a saturated solution.

- ☒ a. The total amount of dissolved solute remains constant.
☒ b. The total mass of undissolved crystals remains constant.
☒ c. When the rate of solvation equals the rate of crystallization, a state of dynamic equilibrium exists.
d. If more solute were added to the container, the total amount of dissolved solute would increase.

8. If two liquids dissolve each other, they are said to be miscible.

9. Look at Figure 16.3 on page 473. Why does the oil float on the vinegar?

Vinegar is water-based, and oil and water are immiscible. The less dense oil floats on the water-based solution.

► Factors Affecting Solubility (pages 474–477)

10. Is the following sentence true or false? The solubility of sodium chloride in water increases to 39.2 g per 100 g of water at 100°C from 36.2 g per 100 g of water at 25°C. true
11. Circle the letter of the sentence that best answers the following question. How does the solubility of solid substances change as the temperature of the solvent increases?
- a. The solubility increases for all solids.
☒ b. The solubility increases for most solids.
c. The solubility remains constant.
12. Look at Table 16.1 on page 474. Which solid substance listed in the table is nearly insoluble at any temperature? barium sulfate
13. How does the solubility of a gas change with an increase in temperature?
As the temperature increases, the solubility of a gas decreases.
14. The directly proportional relationship between the solubility of a gas in a liquid and the pressure of the gas above the liquid is known as Henry's law.

Chapter 16 Solutions Worksheet Answers

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