

14

THE BEHAVIOR OF GASES

SECTION 14.1 PROPERTIES OF GASES (pages 413–417)

This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature.

► Compressibility (pages 413–414)

1. Look at Figure 14.1 on page 413. How does an automobile air bag protect the crash dummy from being broken as a result of impact?

The gases used to inflate the airbag are able to absorb a considerable amount of energy when they are compressed.

2. What theory explains the behavior of gases? kinetic theory

3. Circle the letter next to each sentence that is true concerning the compressibility of gases.

- ☒ a. The large relative distances between particles in a gas means that there is considerable empty space between the particles.
- ☒ b. The assumption that particles in a gas are relatively far apart explains gas compressibility.
- ☒ c. Compressibility is a measure of how much the volume of matter decreases under pressure.
- d. Energy is released by a gas when it is compressed.

► Factors Affecting Gas Pressure (pages 414–417)

4. List the name, the symbol, and a common unit for the four variables that are generally used to describe the characteristics of a gas.

- a. pressure, P , kilopascals
- b. volume, V , liters
- c. temperature, T , kelvins
- d. amount of gas, n , moles

5. What keeps the raft in Figure 14.3 inflated?

The air pressure exerted by the enclosed gas keeps the raft inflated.

Chapter 14 Gases Study Guide Answer Key

Mary Jo Boyer



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Table of Contents Chapter 14 Gases Study Guide Answer Key

1. Understanding the eBook Chapter 14 Gases Study Guide Answer Key
 - The Rise of Digital Reading Chapter 14 Gases Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 14 Gases Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 14 Gases Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 14 Gases Study Guide Answer Key
 - Personalized Recommendations
 - Chapter 14 Gases Study Guide Answer Key User Reviews and Ratings
 - Chapter 14 Gases Study Guide Answer Key and Bestseller Lists
5. Accessing Chapter 14 Gases Study Guide Answer Key Free and Paid eBooks
 - Chapter 14 Gases Study Guide Answer Key Public Domain eBooks
 - Chapter 14 Gases Study Guide Answer Key eBook Subscription Services
 - Chapter 14 Gases Study Guide Answer Key Budget-Friendly Options
6. Navigating Chapter 14 Gases Study Guide Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Chapter 14 Gases Study Guide Answer Key Compatibility with Devices
- Chapter 14 Gases Study Guide Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 14 Gases Study Guide Answer Key
 - Highlighting and Note-Taking Chapter 14 Gases Study Guide Answer Key
 - Interactive Elements Chapter 14 Gases Study Guide Answer Key
- 8. Staying Engaged with Chapter 14 Gases Study Guide Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 14 Gases Study Guide Answer Key
- 9. Balancing eBooks and Physical Books Chapter 14 Gases Study Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 14 Gases Study Guide Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 14 Gases Study Guide Answer Key
 - Setting Reading Goals Chapter 14 Gases Study Guide Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 14 Gases Study Guide Answer Key
 - Fact-Checking eBook Content of Chapter 14 Gases Study Guide Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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