

## Chapter 15 – Acids and Bases

- Acids – have sour taste
  - Vinegar – acetic acid
  - Citrus fruits – citric acid
  - React w metals to produce hydrogen gas
  - React with carbonates and bicarbonates to produce carbon dioxide gas
  - Change the color of certain dyes – green, yellow, blue
- Bases
  - Taste bitter
  - Feel slippery – some soaps
- Bronsted-Lowry
  - Bronsted acid - proton donor
  - Bronsted base – proton acceptor
- Arrhenius theory
  - Acid – anything that produces  $H_3O^+$  in aqueous solution
  - Base – anything that produces  $OH^-$  in aqueous solutions
- When an acid donates proton  $\rightarrow$  conjugate base
- When a base accepts a proton  $\rightarrow$  conjugate acid
- The ion-product constant ( $K_w$ ) is the product of the molar concentrations of  $H^+$  and  $OH^-$  ions at a particular temperature.
- Conjugate acid-base pairs
  - The conjugate base of a strong acid has no measurable strength.
  - $H_3O^+$  is the strongest acid that can exist in aqueous solution.
  - The  $OH^-$  ion is the strongest base that can exist in aqueous solution.
- SLIDE 27 – MEMORIZE
- Diprotic and triprotic acids
  - Diprotic- 2 protons to donate
  - Triprotic- 3 protons to donate
  - May yield more than one hydrogen ion per molecule
  - Ionize in a stepwise manner; they lose one proton at a time
  - An ionization constant expression can be written for each ionization stage
  - Two or more equilibrium constant expressions must often be used to calculate the concentrations of species in the acid solution
- Binary acids –  $(H)_1 \text{ ____}$ 
  - Ex. HF
- Oxoacids –  $(H)_1 \text{ ____}(O)_1$ 
  - Ex.  $HNO_3$
  - \_\_\_\_: any element
  -

# Chapter 15 Acids Bases Study Guide

**Jill Meryl Levy**



## Chapter 15 Acids Bases Study Guide:

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