# Practice Problems: Redox Reactions (Answer Key)

1. Determine the oxidation number of the elements in each of the following compounds:

```
a H<sub>2</sub>CO<sub>3</sub>
H: +1, O: -2, C: +4
b N<sub>2</sub>
N: 0
c Zn(OH)<sub>4</sub><sup>2</sup>·
Zn: 2+, H: +1, O: -2
d NO<sub>2</sub>·
N: +3, O: -2
e LiH
Li: +1, H: -1
f Fe<sub>3</sub>O<sub>4</sub>
Fe: +8/3, O: -2
```

2. Identify the species being oxidized and reduced in each of the following reactions:

 Would you use an oxidizing agent or reducing agent in order for the following reactions to occur?

# **Chapter 19 Practice Problems Redox Reactions**

**Christian G. Meyer** 

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