

Chapter 19- Acids, Bases and Salts. GA chemistry standards:
 SC7.b: compare contrast and evaluate the nature of acids and bases
 SC7.b.1: Arrhenius, Bronsted-Lowery acid/bases
 SC7.b.2: strong vs. weak acids/bases in terms of percent dissociation
 SC7.b.3: hydronium ion concentration (hydrogen ion)
 SC7.b.4: Acid-Base neutralization

19.1 Acid-Base Theories

- Properties of Acids
 - Sour taste
 - React with metals to produce hydrogen gas
 - Aqueous solutions conduct electricity and are call electrolytes
 - Will change the color of acid base indicator
 - Low pH (below 7)
- Properties of Bases
 - Bitter taste
 - Feel slippery
 - Aqueous solutions conduct electricity and are call electrolytes
 - Will change the color of acid base indicator
 - High pH (above 7)
- Arrhenius Acids are defined as hydrogen-containing compounds that ionize to yield hydrogen ions (H^+) in aqueous solutions
 - Monoprotic acid are acids that contain ONE ionizable hydrogen. Examples: HCl or HNO_3
 - Diprotic acid are acids that contain TWO ionizable hydrogen. Examples: H_2S or H_2SO_4
 - Triprotic acids are acids that contain THREE ionizable hydrogen. Examples: H_3PO_4
- Arrhenius bases are defined as compounds that ionize to yield hydroxide ions (OH^-) in aqueous solutions
- Bronsted-Lowery defines an acid as a Hydrogen-ion donor
- Bronsted-Lowery defines a base as a hydrogen-ion acceptor
 - Conjugate acid is the particle formed when a base gains a hydrogen ion
 - Conjugate base is the particle formed when an acid loses a hydrogen ion
 - Conjugate acid-base pair consist of two substances that are related by the loss or gain of a single hydrogen ion



- Conjugate acid-base pairs: NH_3 & NH_4^+ , H_2O & OH^-
- If the compounds differ by more than one hydrogen ion (or any other element) they CAN NOT be classified as conjugate acid-base pairs.
- Lewis acid is a substance that can accept a pair of electrons to form a covalent bond
- Lewis base is a substance that can donate a pair of electrons to form a covalent bond

Type	Acid	Base
Arrhenius	H^+ producer (donor)	OH^- producer
Bronsted-Lowery	H^+ producer (donor)	H^+ acceptor
Lewis	Electron-pair acceptor	Electron-pair donor

19.2 Hydrogen Ions and Acidity

- Self-ionization is where water molecules production ions

$$\text{H}_2\text{O}(l) \rightleftharpoons \text{H}^+(aq) + \text{OH}^-(aq)$$

Hydrogen ion Hydroxide ion
- Neutral solution is where the $[H^+]$ and $[OH^-]$ are equal ([] = concentration of)

Chapter 19 Acids And Bases Study Guide For Content Mastery

Tobias Bleicker



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