

12**STOICHIOMETRY****Chapter Test A****A. Matching**

Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line.

Column A**Column B**

- | | |
|----------------------------|--|
| _____ 1. actual yield | a. the ratio of the actual yield to the theoretical yield, expressed as a percentage |
| _____ 2. limiting reagent | b. the amount of product formed when a reaction is carried out in the laboratory |
| _____ 3. theoretical yield | c. the reactant that determines the amount of product that can be formed in a reaction |
| _____ 4. stoichiometry | d. the reactant that is not completely used up in a reaction |
| _____ 5. percent yield | e. the calculated amount of product that might be formed during a reaction |
| _____ 6. excess reagent | f. the calculation of quantities in chemical equations |

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 7. Which of these expressions is an incorrect interpretation of the balanced equation?
- $$2S(s) + 3O_2(g) \rightarrow 2SO_2(g)$$
- a. 2 atoms S + 3 molecules $O_2 \rightarrow$ 2 molecules SO_2
b. 2 g S + 3 g $O_2 \rightarrow$ 2 g SO_2
c. 2 mol S + 3 mol $O_2 \rightarrow$ 2 mol SO_2
d. none of the above
- _____ 8. In a chemical reaction, the mass of the products
- a. is less than the mass of the reactants.
b. is greater than the mass of the reactants.
c. is equal to the mass of the reactants.
d. has no relationship to the mass of the reactants.
- _____ 9. How many liters of oxygen are required to react completely with 1.2 liters of hydrogen to form water?
- $$2H_2(g) + O_2(g) \rightarrow 2H_2O(g)$$
- a. 1.2 L
b. 0.6 L
c. 2.4 L
d. 4.8 L

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