Chapter 13: Properties of Solutions

Answers to selected problems:

23. a) unsaturated b) saturated

24. a) 25 g

33. Solubility of Helium: 5.6 x 10 4 M; Solubility of N2 = 9.0 x 10 4 M.

34. Solubility of Oxygen: 2.5 x 104 M. Hint: Be sure to change torr to atm.

35. a) 2.15 %; b) 3.15 ppm

38. a) Mole fraction of phenol: 0.0285

b) 5.66 % c) 0.638 m

40. a) 0.125 M Aluminum Sulfate b) 0.140 M

c) 0.63 M (hint: use MV = MV, see Chapter 4)

41. a) 4.70 m b) 0.235 m

43. **To make this problem easier, I would assume 1,000 L or 1000 mL.

a) 43.01%

b) mole fraction = 0.122

c) 7.69 m

d) 5.828 Mi

46. a) 0.0439 b) 0.498 m c) 0.417 M

(Hint: on this problem, you will have to use the density of thiopene to get the mL of thiopene and you will also have to you use the density of toluene to get the mass(g) of toluene.

 a) 0.278 mol
b) 6.25 x 10⁻⁵ moles (assume the mass of the solvent is approx. equal to the mass of the solution for this problem)

c) 3.29 x 10⁻³ mol

** Assume you have 1.000L or 1000 mL. You can also use density to get the total grams.
 Answer: 71%

Chapter 13 Review Solutions Section 2 Answers

William Alexander Hannah, Akhil Behl

Chapter 13 Review Solutions Section 2 Answers:

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