

## Section Review

## Objectives

- Describe how the amounts of reactants and products change in a chemical system at equilibrium
- Identify three stresses that can change the equilibrium position of a chemical system
- Explain what the value of  $K_{eq}$  indicates about the position of equilibrium

## Vocabulary

- reversible reaction
- chemical equilibrium
- equilibrium position
- Le Châtelier's principle
- equilibrium constant ( $K_{eq}$ )

## Key Equation

$$K_{eq} = \frac{[C]^c \times [D]^d}{[A]^a \times [B]^b}$$

When  $aA + bB \rightleftharpoons cC + dD$

## Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

1. A reversible reaction is one in which the reaction can proceed in both directions.

2. A reversible reaction is one in which the reaction can proceed in both directions.

3. A reversible reaction is one in which the reaction can proceed in both directions.

4. A reversible reaction is one in which the reaction can proceed in both directions.

5. A reversible reaction is one in which the reaction can proceed in both directions.

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13. A reversible reaction is one in which the reaction can proceed in both directions.

14. A reversible reaction is one in which the reaction can proceed in both directions.

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16. A reversible reaction is one in which the reaction can proceed in both directions.

17. A reversible reaction is one in which the reaction can proceed in both directions.

18. A reversible reaction is one in which the reaction can proceed in both directions.

19. A reversible reaction is one in which the reaction can proceed in both directions.

20. A reversible reaction is one in which the reaction can proceed in both directions.

# Chapter 18 2 Reversible Reactions And Equilibrium

## Worksheet Answers

**JG Myers**



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