Chapter 19- Acids, Bases and Salts. GA chemistry standards: SC7.b: compare contrast and evaluate the nature of acids and bases SC7.b.1: Arrhenius, Bronsted-Lowery acid/bases SC7.b.2: strong vs. weak acids/bases in terms of percent dissociation SC7.b.3: hydronium ion concentration (hydrogen ion) SC7.b.4: Acid-Base neutralization

- 19.1 Acid-Base Theories
 - Properties of Acids
 - o Sour taste
 - React with metals to produce hydrogen gas
 - Aqueous solutions conduct electricity and are call electrolytes.
 - Will change the color of acid base indicator.
 - Law pH (below 7)
 - Properties of Bases
 - o Bitter taste
 - Feel slippery
 - Aqueous solutions conduct electricity and are call electrolytes
 - Will change the color of acid base indicator
 - o High pH (above 7)
 - Arrhenius Acids are defined as hydrogen-containing compounds that ionize to yield hydrogen ions (H**) in aqueous solutions
 - Monoprotic acid are acids that contain ONE ionizable hydrogen. Examples: HCl or HNO.
 - Diprotic acid are acids that contain TWO ionizable hydrogen. Examples: H₂S or H₂SO_x
 - Triprotic acids are acids that contain THREE ionizable hydrogen. Examples: H₂PO₄
 - Arrhenius bases are defined as compounds that ionize to yield hydroxide ions (OH*) in aqueous solutions
 - Bronsted-Lowery defines an acid as a Hydrogen-lon donor
 - Bronsted-Lowery defines a base as a hydrogen-ion acceptor
 - Conjugate acid is the particle formed when a base gains a hydrogen ion
 - Conjugate base is the particle formed when an acid looses a hydrogen ion
 - Conjugate acid-base pair consist of two substances that are related by the loss or gain of a single hydrogen ion

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- Conjugate acid-base pairs: NH, &NH, *, H,O & OH*
- If the compounds differ by more that one hydrogen ion (or any other element) they CAN NOT be classified as conjugate acid-base pairs.
- Lewis acid is a substance that can accept a pair of electrons to form a covalent bond.
- Lewis base is a substance that can donate a pair of electrons to form a covalent bond.

Туре	Acid	Base
Arrhenius	H" producer (donor)	OH producer
Bronsted-Lowery	H [*] producer (donor)	H* acceptor
Lewis	Electron-pair acceptor	Electon-pair donor

- 19.2 Hydrogen lons and Acidity
 - Self-ionization is where water molecules production ions.

$$H_2O(I) \rightleftharpoons H^+(aq) + OH^-(aq)$$

- Hydrogen ion Hydroxide ion
- Neutral solution is where the [H**] and [OH*] are equal ([]-concentration of)

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Darrell D. Ebbing, Susan M. Schelble

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