

Question Number	Acceptable Answers	Mark
18(c)(i)	<p>IGNORE sf except 1</p> <p>If answer is 3.84 (mol dm⁻³), award 3 marks</p> <p>If not, number of moles of acid = $\frac{24.0 \times 4}{1000} = 0.096$ (1)</p> <p>EITHER number of moles ammonia = 0.096 in 25 cm³ (1)</p> <p>concentration of ammonia $= \frac{0.096 \times 1000}{25}$ $= 3.84 \text{ (mol dm}^{-3}\text{)}$ (1)</p> <p>OR number of moles ammonia = 0.288 in 75 cm³ (1)</p> <p>concentration of ammonia $= \frac{0.288 \times 1000}{75}$ $= 3.84 \text{ (mol dm}^{-3}\text{)}$ (1)</p> <p>IGNORE unit unless incorrect</p> <p>ALLOW ecf in both methods on their number of moles of ammonia</p>	3

Question Number	Acceptable Answers	Mark
18(c)(ii)	<p>IGNORE sf except 1 (concentration of ammonia in trichloromethane =) 0.16 (mol dm⁻³)</p> <p>ALLOW ecf from their answer to (c)(i), provided it is less than 4.0 and given to 2 or more sf</p>	1

Question Number	Acceptable Answers	Reject	Mark
18(c)(iii)	<p>Expression for K_c and answer needed for the mark</p> $K_c = \frac{[\text{NH}_3(\text{aq})]}{[\text{NH}_3(\text{CHCl}_3)]}$ <p>ALLOW one state symbol missing</p> $= \frac{3.84}{0.16}$ $= 24(.0)$ <p>IGNORE sf, including 1 sf, and units</p> <p>ALLOW ecf from answers to (c)(i) and (c)(ii)</p>	<p>K_c expressions without both state symbols</p>	1

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