

7. GROUPS 1 & 2

SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

- 1 Magnesium oxide, known as magnesia, can be produced by the thermal decomposition of magnesium carbonate.

The equation for the reaction is



- *(a) Other Group 2 carbonates also undergo thermal decomposition.

State and explain the trend in thermal stability of carbonates down Group 2.

(4)

Thermal stability increases down the group

Because ionic radius increases as you go down the group

This results in less polarisation (distortion) of the carbonate ion

And less weakening of the covalent bonds (C-O) within the carbonate ion

- *(c) State why carrying out a flame test is not an appropriate way to confirm the presence of magnesium ions in the sample of magnesia.

(1)

Magnesium ions give no flame colour

OR

Energy is emitted with Mg ions is outside of the visible region

(Total for Question 1 = 15 marks)

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Table of Contents Edexcel Igcse Biology Past Question Papers

1. Understanding the eBook Edexcel Igcse Biology Past Question Papers
 - The Rise of Digital Reading Edexcel Igcse Biology Past Question Papers
 - Advantages of eBooks Over Traditional Books
2. Identifying Edexcel Igcse Biology Past Question Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Edexcel Igcse Biology Past Question Papers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Edexcel Igcse Biology Past Question Papers
 - Personalized Recommendations
 - Edexcel Igcse Biology Past Question Papers User Reviews and Ratings
 - Edexcel Igcse Biology Past Question Papers and Bestseller Lists
5. Accessing Edexcel Igcse Biology Past Question Papers Free and Paid eBooks
 - Edexcel Igcse Biology Past Question Papers Public Domain eBooks
 - Edexcel Igcse Biology Past Question Papers eBook Subscription Services
 - Edexcel Igcse Biology Past Question Papers Budget-Friendly Options
6. Navigating Edexcel Igcse Biology Past Question Papers eBook Formats
 - ePub, PDF, MOBI, and More
 - Edexcel Igcse Biology Past Question Papers Compatibility with Devices
 - Edexcel Igcse Biology Past Question Papers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Edexcel Igcse Biology Past Question Papers
 - Highlighting and Note-Taking Edexcel Igcse Biology Past Question Papers
 - Interactive Elements Edexcel Igcse Biology Past Question Papers
8. Staying Engaged with Edexcel Igcse Biology Past Question Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Edexcel Igcse Biology Past Question Papers
9. Balancing eBooks and Physical Books Edexcel Igcse Biology Past Question Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Edexcel Igcse Biology Past Question Papers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Edexcel Igcse Biology Past Question Papers
 - Setting Reading Goals Edexcel Igcse Biology Past Question Papers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Edexcel Igcse Biology Past Question Papers
 - Fact-Checking eBook Content of Edexcel Igcse Biology Past Question Papers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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