

7. GROUPS 1 & 2

SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

- 1 Magnesium oxide, known as magnesia, can be produced by the thermal decomposition of magnesium carbonate.

The equation for the reaction is



- *(a) Other Group 2 carbonates also undergo thermal decomposition.

State and explain the trend in thermal stability of carbonates down Group 2.

(4)

Thermal stability increases down the group

Because ionic radius increases as you go down the group

This results in less polarisation (distortion) of the carbonate ion

And less weakening of the covalent bonds (C-O) within the carbonate ion

- *(c) State why carrying out a flame test is not an appropriate way to confirm the presence of magnesium ions in the sample of magnesia.

(1)

Magnesium ions give no flame colour

OR

Energy is emitted with Mg ions is outside of the visible region

(Total for Question 1 = 15 marks)

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RJ Shavelson



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