

Mark Scheme (Results)

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# Edexcel Mathematics C1 May 2014 Mark Scheme

**Sue Johnston-Wilder, Clare Lee, David  
Pimm**



## **Edexcel Mathematics C1 May 2014 Mark Scheme:**

*Learning to Teach Mathematics in the Secondary School* Sue Johnston-Wilder, Clare Lee, David Pimm, 2016-09-13

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 $0 5 3 5 0 1429$

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40 g of  $H_2O$  molar mass of  $H_2O$  is 18gm mole and molar mass of  $CH_3OH$  is 32gm mole solution moles of  $CH_3OH$  4 5 32 0 14  
mole moles of  $H_2O$  40 18 2 2 moles therefore according to the equation

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time the total pressure  $p_a \times a \times t$  mole fraction  $a$  left frac mole  $a$  moles total right sum  $n_1$  all  $x$   $n_1$  mole fraction of substance  
 $a$  is often denoted by  $x_a$  and

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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