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Edexcel M1 Maths Bronze Paper

Keith Pledger



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 $2S + 3O_2 \rightarrow 2SO_3$ $\Delta H = -7914 \text{ kJ}$ is this endothermic or exothermic?

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 $2H_2(g) + O_2(g) \rightarrow 2H_2O(l)$ energy endothermic the transfer of heat from the surroundings to the system energy
 $2Hg(s) + 2Hg(l) \rightarrow 2Hg_2(g)$

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 $35.5 \text{ g} \times 1 \text{ mole KCN} / 0.5453 \text{ moles} = 65.1 \text{ g}$
 $0.5453 \text{ mole moles} = 5$ when a 15 g sample of $SnCl_2$ dissolves in 80 g of water in a calorimeter the temperature drops from 42°C to 15°C calculate the ΔH for the process
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