

# Chemistry Unit 1 Review Sheet



## The History of the Atom

You don't need to memorize any specifics about the history or the scientists. However, you may be asked to look at some information about a scientist and/or an experiment and identify what major discovery was made. For example, JJ Thomson's cathode ray tube experiment helped identify the Electron using the idea of Charge attraction and repulsion

## Vocabulary

Matter

Atom

Nucleus

Proton

Neutron

Electron

Element

Atomic Number

Mass Number

Ion

Net Charge

AMU

Orbital

Wave Mechanical Model

Bohr Model

Electron Configuration

Ground State

Excited State

Bright Line Spectrum

Isotope

Atomic Mass

$$\text{Mass Number} = p^+ + n^0$$

\*Electrons DON'T Count!

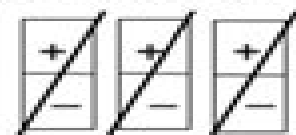
Nuclear      Mass #  $\rightarrow$  12  
Symbols      Atomic #  $\rightarrow$  6      **C**

Proton	1 amu
Neutron	1 amu
Electron	0 amu

## Ions and Net Charge



Neutral Atoms have EQUAL  $p^+$  &  $e^-$



( $3p^+$ ) & ( $3e^-$ )  
( $+3$ ) + ( $-3$ )  
 $3 - 3 = 0$   
**Li<sup>0</sup>**

IONS DO NOT have same #  $p^+$  &  $e^-$



( $3p^+$ ) & ( $2e^-$ )  
( $+3$ ) + ( $-2$ )  
 $3 - 2 = +1$   
**Li<sup>+1</sup>**

\*\*Positive Charge  $\rightarrow$  More  $p^{++}$   
\*\*Negative Charge  $\rightarrow$  More  $e^{--}$

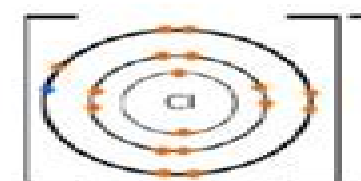
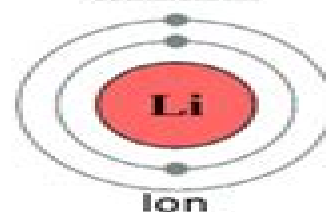
## Bohr Models

1<sup>st</sup> Shell = 2  $e^-$ s

2<sup>nd</sup> Shell = 8  $e^-$ s

3<sup>rd</sup> Shell = \*8\*  $e^-$ s

Neutral



## Electron Configurations

Neutral Ground State  $\rightarrow$  Ca: 2 - 8 - 8 - 2

Ionic Ground State  $\rightarrow$  Ca<sup>2+</sup>: 2 - 8 - 8

Excited State  $\rightarrow$  Ca: 2 - 8 - 7 - 3

## Excited Electrons

- Electrons closer to the nucleus have LOWER energy.
- Electrons further from the nucleus have HIGHER energy.
- An Excited Electrons GAINS energy and Jumps to a higher energy level.
- A De-Excited Electrons RELEASES energy and falls back to ground state.

$\rightarrow$  THIS IS WHAT PRODUCES THE BRIGHT LINE SPECTRUM



# Chemistry Unit 1review Answers

**Ensheng Dong**



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**Table of Contents Chemistry Unit 1review Answers**

1. Understanding the eBook Chemistry Unit 1review Answers
  - The Rise of Digital Reading Chemistry Unit 1review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 1review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Unit 1review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 1review Answers
  - Personalized Recommendations
  - Chemistry Unit 1review Answers User Reviews and Ratings
  - Chemistry Unit 1review Answers and Bestseller Lists
5. Accessing Chemistry Unit 1review Answers Free and Paid eBooks
  - Chemistry Unit 1review Answers Public Domain eBooks
  - Chemistry Unit 1review Answers eBook Subscription Services
  - Chemistry Unit 1review Answers Budget-Friendly Options
6. Navigating Chemistry Unit 1review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Unit 1review Answers Compatibility with Devices
  - Chemistry Unit 1review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Unit 1review Answers
  - Highlighting and Note-Taking Chemistry Unit 1review Answers
  - Interactive Elements Chemistry Unit 1review Answers
8. Staying Engaged with Chemistry Unit 1review Answers

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Unit 1review Answers
9. Balancing eBooks and Physical Books Chemistry Unit 1review Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Unit 1review Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Unit 1review Answers
- Setting Reading Goals Chemistry Unit 1review Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Unit 1review Answers
- Fact-Checking eBook Content of Chemistry Unit 1review Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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