

Phase Change

- **Deposition**-gas to a solid; **Sublimation**-solid to gas; **Vaporization**- liquid to gas; **Condensation**-gas to liquid
- **Physical change** = melting, freezing, dissolving, evaporating;
- **Chemical change** = boiling an egg, converting, exploding, burning, digesting.

Atomic Numbers

- Number that identifies an element-**Atomic number**
- Weighted average mass of all occurring isotopes of an element-**Atomic mass**
- Symbolic method for expressing the composition of an atomic nucleus-**Atomic notation**

Electron Configurations

- Shorthand description of electrons by sublevels according to increasing energy-**Electron configuration**
- **Sublevels:** 2,6,10,14 electrons in the sublevels; **1s 2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d 7p;**



Chemical Formulas

- Name of a chemical compound that indicates the # of atoms of each element-**Chemical formula**
- **Number of atoms** from formula- $\rightarrow \text{C}_{17}\text{H}_{19}\text{NO}_3 = 40\text{atoms}$

NaHCO₃=sodium hydrogen carbonate; **Mg(ClO₄)₂**= Magnesium Perchlorate; **PbO₂**=Lead (IV) oxide;

Hg₂Br₂=Mercury (II) bromide; **Al₂(CrO₄)₃**=Aluminum Chromate; **ZnCl₂**= Zinc Chloride; Nitrous acid=**HNO₂**;

Hypochlorous acid=**HClO**; hydrochloric=**HCl**; **H₂S**=hydrosulfuric acid; **HClO₄**=perchloric acid; **H₂SO₄**=sulfuric acid; **HNO₃**=nitric acid; **SnO₂**=Tin dioxide; sulfuric acid=**H₂SO₄**;

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James E. Brady



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