

DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Solve the problems below

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1. The empirical formula of a compound is NO_2 , its molecular mass is 92 g/mol. What is its molecular formula?

2. The empirical formula of a compound is CH_2 , its molecular mass is 70 g/mol. What is its molecular formula?

3. A compound is found to be 40.0% carbon, 6.7% hydrogen, and 53.5% oxygen. Its molecular mass is 60 g/mol. What is its molecular formula?

4. A compound is 64.9% carbon, 13.5% hydrogen, and 21.6% oxygen. Its molecular mass is 74 g/mol. What is its molecular formula?

5. A compound is 54.5% carbon, 9.1% hydrogen, and 36.4% oxygen. Its molecular mass is 88 g/mol. What is its molecular formula?

Determining Molecular Formulas True Formulas Answers

Ying-Ying Zheng



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