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2.1 (a) (F/P,10%,20) = 6.7275 (b) (A/F,4%,8) = 0.10853(c) (P/A,8%,20) = 9.8181(d) (A/P,20%,28) = 0.20122(e) $(F/A_30\%, 15) = 167.2863$ 2.2 P = 30,000(P/F,10%,8)=30,000(0.4665)= \$13,995 2.3 F = 15,000(F/P,6%,25)= 15.000(4.2919)= 64.378.502.4 (a) F = 885,000 + 100,000(F/P,10%,3)= 885,000 + 100,000(1.3310)= \$1,018,000(b) Spreadsheet function is = -FV(10%,3,,100000) + 885000 Display is \$1,018,000 2.5 (a) P = 19,000(P/F,10%,7)= 19,000(0.5132)= \$9750.80(b) If the calculator function is PV(10,7,0,19000), display is P = \$-9750,00 (c) If the spreadsheet function is = -PV(10%,7,,19000), display is \$9750.00 2.6 (a) Total for 7 lots is 7(120,000) = \$840,000

> P = 840,000(P/F,10%,2) = 840,000(0.8264)

= 5694.176

(b) If the calculator function is PV(10,2,0,840000), display is P = \$-694,214.88

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web molarity has units of dfrac text mol text liter litermol which can be abbreviated as molar or text m m pronounced molar the molar concentration of the solute is sometimes abbreviated by putting square brackets around the chemical formula of **molality definition formula difference between molarity molality** - Apr 30 2022

web molarity is the number of moles of a substance per litre of solution also known as molar concentration a capital m signifies solutions labelled with molar concentration a 1 0 m solution contains 1 mole of solvent per litre of solution molality is the number of solvent moles per kilogram

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web 1 molarity mv mass molar mass x $0.4500 \, l\, 0.825 \, g\, 141\, 9579 \, g\, mol \, x\, 0.0129 \, m\, 2$ molality $0.825 \, g\, 141\, 9579 \, g\, mol \, 0.00581158 \, mol \, 0.00$

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web apr 7 2014 molarity moles of solute liters of solution 8 4 2 2 a first convert 250 ml to liters 250 1000 0 25 then calculate molarity 5 moles 0 25 liters 20 m 3 c a solution with molarity 2 requires 2 m of n a oh per liter so $4 \times 2 \times 8 = 4 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 8 = 4 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 8 = 4 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 8 = 4 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 8 = 4 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 8 = 4 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 8 = 4 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 8 = 4 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 8 = 4 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so $4 \times 2 \times 10^{-2}$ m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so 4×10^{-2} m of n a oh per liter so

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web answer the correct option is c explanation the molarity of a mixture mmix can be calculated using the following formula m mix m 1 v 1 m 2 v 2 v 1 v 2 1 5 x 480 1 2 x 520 480 520 720 624 1000 1 344 m q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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web 1 calculate the molarity of a solution which contains 0 40 mol of a substance dissolved in 1 6 l of a solution 2 what is the molarity of a solution containing 325 g of nacl dissolved in 750 ml of solution 3 140 g of kcl is dissolved in 600 ml of water what is the molarity 4 724 4 g of ammonium phosphate in 4500 ml of alcohol

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web determine the molarity for each of the following solutions 0 444 mol of cocl 2 in 0 654 l of solution 98 0 g of phosphoric acid h 3 po 4 in 1 00 l of solution 0 2074 g of calcium hydroxide ca oh 2 in 40 00 ml of solution 10 5 kg 16 11 molality chemistry libretexts - Aug 15 2023

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