Chapter 2: Elements are the Building Blocks of Matter Study Guide "Without pressure, the work doesn't get done at all"

William Sarovan (1902-1921)

Steps for studying for the Ch 2 test

- Re-sead your notes and use your Cornell-questions to quiz yourself.
- Re-read pages 40 73 of your textbook.
- Be able to define the key vocabulary from the chapter.
- Review your Reading Checks and Work Book, making sure all is complete.
- Be able to answer the questions below. They are a guide for your studying.

metalloids

Complete the online quizzes at http://www.bcscience.com/bc9/

Chapter 2 VOCABULARY (pages 40 -73)**Make vocab cards and quiz yourself if you need to

elements		atomic		period	electron
compounds		number	•	chemical	shells
chemical	*	atomic mass		family	valence
symbol		atomic mass	•	alkali metals	electrons
chemical		umit	•	allcaline	valence shell
properties	•	ion charge		earth metals	noble gas
metals		ion	•	halogens	stability
non-metals		multiple ion	•	noble gases	
periodic		charge		Bohr model	

of the atom.

Section 2.1: Elements

taib le

- How many elements are there?
- Give 2 examples of elements at each of the following states:

a) solid b) liquid c) gas

- 3. Most of the elements names are derived from what languages?
- Be familiar with the physical and chemical properties, some uses, and interesting information of the following elements.

a) Hydrogen e) Chlorine
b) Iron f) Mercury
c) Oxygen g) Silver
d) Sodium h) Silicon

5. Complete the Section 2.1 Quiz online at http://www.bcscience.com/bc9/

Section 2.2: The Periodic Table and Chemical Properties

- a) Who is Dmitn Mendeleev?
 - b) Describe his early work on organizing the elements.
 - c) Describe Mendeleev's special insight in the organization of elements.
- What does the Periodic Table show for each element?
- 3. a). Differentiate atomic number from atomic mass and choose a specific element for example.
 - b). Describe the pattern of each on the Penodic Table.
- Differentiate atoms from ions.
- a) What are the 2 types of ions?

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Table of Contents Bc Science 9 Study Guide

- 1. Understanding the eBook Bc Science 9 Study Guide
 - The Rise of Digital Reading Bc Science 9 Study Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Bc Science 9 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Bc Science 9 Study Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Bc Science 9 Study Guide

- Personalized Recommendations
- Bc Science 9 Study Guide User Reviews and Ratings
- Bc Science 9 Study Guide and Bestseller Lists
- 5. Accessing Bc Science 9 Study Guide Free and Paid eBooks
 - Bc Science 9 Study Guide Public Domain eBooks
 - Bc Science 9 Study Guide eBook Subscription Services
 - Bc Science 9 Study Guide Budget-Friendly Options
- 6. Navigating Bc Science 9 Study Guide eBook Formats
 - o ePub, PDF, MOBI, and More
 - Bc Science 9 Study Guide Compatibility with Devices
 - Bc Science 9 Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Bc Science 9 Study Guide
 - Highlighting and Note-Taking Bc Science 9 Study Guide
 - Interactive Elements Bc Science 9 Study Guide
- 8. Staying Engaged with Bc Science 9 Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Bc Science 9 Study Guide
- 9. Balancing eBooks and Physical Books Bc Science 9 Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Bc Science 9 Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Bc Science 9 Study Guide
 - Setting Reading Goals Bc Science 9 Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Bc Science 9 Study Guide

- Fact-Checking eBook Content of Bc Science 9 Study Guide
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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