Objective 4: Students will know the different types of chemical reactions by looking at a chemical equation.

- 1. Explain what happens during each type of chemical reaction.
  - Synthesis- two or more compounds become one
    - Combustion- oxygen is a reactant, water and carbon dioxide are products
    - Decomposition- one compound becomes two or more
    - Single displacement- one switch occurs
    - Double displacement- two switches occur
- 2. State what type of chemical reaction each of the following equations show.
  - C<sub>10</sub>H<sub>8</sub> + 12 O<sub>2</sub> ---> 10 CO<sub>2</sub> + 4 H<sub>2</sub>O

#### combustion

8 Fe + S<sub>8</sub> ---> 8 FeS

# synthesis

2 H<sub>2</sub>O ---> 2 H<sub>2</sub> + O<sub>2</sub>

## decomposition

- Mg + 2 H<sub>2</sub>O ---> Mg(OH)<sub>2</sub> + H<sub>2</sub>
  - single displacement
- Pb(NO<sub>3</sub>)<sub>2</sub> + 2 KI ---> Pbl<sub>2</sub> + 2 KNO<sub>3</sub>

## double displacement

NaOH + KNO<sub>3</sub> --> NaNO<sub>3</sub> + KOH

#### double displacement

CH<sub>4</sub> + 2 O<sub>2</sub> --> CO<sub>2</sub> + 2 H<sub>2</sub>O

#### combustion

2 Fe + 6 NaBr --> 2 FeBr<sub>3</sub> + 6 Na

## single displacement

CaSO<sub>4</sub> + Mg(OH)<sub>2</sub> --> Ca(OH)<sub>2</sub> + MgSO<sub>4</sub>

## double displacement

Pb + O2 --> PbO2

- synthesis
- Na<sub>2</sub>CO<sub>3</sub> --> Na<sub>2</sub>O + CO<sub>2</sub>
   decomposition

# **Describing Chemical Reactions Work Sheet With Answers**

**Christian G. Meyer** 

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