Pre-Lab for Determination of Solubility Product Constant

 Write the solubility product constant (K_{sp}) for the following reaction as a function of the concentrations of the products:

$$Ca(IO_3)_2(s)$$
 $Ca^{2+}(aq) + 2IO_3^{-}(aq)$

- Ca(IO₃)₂ will ionize in water to produce IO₃⁻ ions. The IO₃⁻ ions will react with KI. Write the reaction for this reaction.
- 3. We will be using starch as an indicator. Why?
- 4. In this experiment, we will first produce I₂ by mixing calcium iodate with excess amount of KI. Based on the following reactions, how many moles of I₂ will be produced if we started with 6 moles of calcium iodate and excess amount of KI?

$$Ca(IO_3)_2(s)$$
 \longrightarrow $Ca^{2+}(aq) + 2IO_3^-(aq)$
 $IO_3^-(aq) + 5I^-(aq) + 6H^+(aq) \longrightarrow 3I_2(aq) + 3H_2O(l)$

5. In this experiment, we will titrate I₂ with Na₂S₂O₃ (sodium thiosulfate). The reaction is provided below. If we needed 22.5 mL of 0.15 M Na₂S₂O₃ to reach the endpoint, how many moles of I₂ were present in the solution? (Must show all calculation steps.)

$$I_2 (aq) + 2S_2O_3^{2-}(aq) \longrightarrow 2I^-(aq) + S_4O_6^{2-}(aq)$$

- How many moles of Ca(IO₃)₂ are needed to produce the amount of I₂ in Problem #5?
- 7. The solubility product constant (K_{sp}) for calcium iodate is 6.47 x 10⁻⁶. The molar mass of calcium iodate is 389.88 g/mol. 175.4 g of calcium iodate is added to 1 liter of water. What would be the concentrations of Ca²⁺ and IO₃ at equilibrium? Would this solution form a precipitate? Is the solution unsaturated, saturated, or supersaturated when it reaches equilibrium? (Must show all calculation steps.)

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