

# **DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)**

Name: \_\_\_\_\_

Solve the problems below:

1. The empirical formula of a compound is  $\text{PCl}_2$ . Its molecular mass is 263 g/mol. What is its molecular formula?

Solution given by student is:

Answer

2. The empirical formula of a compound is  $\text{CH}_2$ . Its molecular mass is 78 g/mol. What is its molecular formula?

Solution given by student is:

Answer

3. A compound is found to contain 69.8% carbon, 6.7% hydrogen and 23.5% oxygen. Its molecular mass is 174 g/mol. What is its molecular formula?

Solution given by student is:

Solution given by student is:

Solution given by student is:

Answer

4. A compound is found to contain 69.8% carbon, 6.7% hydrogen and 23.5% oxygen. Its molecular mass is 174 g/mol. What is its molecular formula?

Solution given by student is:

Solution given by student is:

Solution given by student is:

Answer

5. A compound is found to contain 69.8% carbon, 6.7% hydrogen and 23.5% oxygen. Its molecular mass is 174 g/mol. What is its molecular formula?

Solution given by student is:

Solution given by student is:

Solution given by student is:

Answer

# Determining Molecular Formulas True Formulas Answer Key

**William Matthai, Creager, Black**



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