#### CHAPTER 14 REVIEW

## Acids and Bases

#### SECTION 1

SHORT ANSWER Answer the following questions in the space provided.

1. Name the following compounds as acids:

sulfuric acid a. H<sub>2</sub>SO<sub>4</sub>

sulfurous acid b. H<sub>2</sub>SO<sub>3</sub>

hydrosulfuric acid e. H<sub>2</sub>S

perchloric acid d. HClO<sub>4</sub>

hydrocyanic acid e. hydrogen cyanide

- 2. H<sub>2</sub>S Which (if any) of the acids mentioned in item 1 are binary acids?
- Write formulas for the following acids:

HNO<sub>2</sub> a. nitrous acid

HBr b, hydrobromic acid

H<sub>3</sub>PO<sub>4</sub> c. phosphoric acid

CH<sub>3</sub>COOH d. acetic acid

HCIO e. hypochlorous acid

Calcium selenate has the formula CaSeO<sub>4</sub>.

H<sub>2</sub>SeO<sub>4</sub> a. What is the formula for selenic acid?

H<sub>2</sub>SeO<sub>3</sub> b. What is the formula for selenous acid?

Use an activity series to identify two metals that will not generate hydrogen gas when treated with an acid.

Choose from Cu, Ag, Au, Pt, Pd, or Hg.

6. Write balanced chemical equations for the following reactions of acids and bases:

a. aluminum metal with dilute nitric acid

 $2AI(s) + 6HNO_3(aq) \rightarrow 2AI(NO_3)_3(aq) + 3H_2(g)$ 

b. calcium hydroxide solution with acetic acid

 $Ca(OH)_2(aq) + 2CH_3COOH(aq) \rightarrow Ca(CH_3COO)_2(aq) + 2H_2O(I)$ 

# **Describing Acids And Bases Section Review**

Julia Burdge

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