

$$.777g \text{ H}_2\text{O} \frac{1 \text{ mol H}_2\text{O}}{18.02g \text{ H}_2\text{O}} = .0431$$

0.0431 moles H_2O

2. Calculate the moles of copper in your sample of copper chloride?

$$.742g \text{ Cu} \frac{1 \text{ mol Cu}}{63.546g \text{ Cu}} = 0.0117$$

0.0117 moles Cu

3. Calculate the moles of chlorine in your sample of copper chloride?

$$M_{\text{Cl}} = M_{\text{Cu}} - M_{\text{hydrate}} = 0.7415 - 0.227 = .515g \text{ Cl}$$

$$.515g \text{ Cl} \frac{1 \text{ mol Cl}}{35.4527} = 0.0145$$

0.0145 moles Cl

4. Write the proper chemical formula for the compound that you tested.

$$\text{Moles Cu} = 0.0117 / 0.0117 = 1.45 = 5$$

$$\text{Moles Cl} = 0.0145 / 0.0117 = 1.245 = 6$$



Correct formula for initial hydrated salt



5. Explain using chemical equations the process of a) dehydrating the hydrate salt b) dissolving the dehydrated salt c) the change in copper observed after adding the aluminum wire to the solution.

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